



# CHEMISTRY

## CHEAT SHEET

### WRITING IONIC FORMULAS

#### IONIC COMPOUNDS CHEAT SHEET

##### WRITING IONIC FORMULAS

**How to Write Ionic Formulas**

1. Write the symbols for the cation and anion, including their charges.

2. Write the cation symbol first, followed by the anion symbol.

3. Use the criss-cross method to determine the subscripts for each ion.

4. Write the formula with the subscripts.

5. Simplify the formula if possible.

Ion	Charge	Symbol
Hydrogen	+1	H <sup>+</sup>
Lithium	+1	Li <sup>+</sup>
Sodium	+1	Na <sup>+</sup>
Potassium	+1	K <sup>+</sup>
Ammonium	+1	NH <sub>4</sub> <sup>+</sup>
Calcium	+2	Ca <sup>2+</sup>
Magnesium	+2	Mg <sup>2+</sup>
Barium	+2	Ba <sup>2+</sup>
Strontium	+2	Sr <sup>2+</sup>
Lead	+2	Pb <sup>2+</sup>
Lead	+4	Pb <sup>4+</sup>
Aluminum	+3	Al <sup>3+</sup>
Iron	+2	Fe <sup>2+</sup>
Iron	+3	Fe <sup>3+</sup>
Copper	+1	Cu <sup>+</sup>
Copper	+2	Cu <sup>2+</sup>
Zinc	+2	Zn <sup>2+</sup>
Nickel	+2	Ni <sup>2+</sup>
Cadmium	+2	Cd <sup>2+</sup>
Mercury	+1	Hg <sup>+</sup>
Mercury	+2	Hg <sup>2+</sup>
Fluoride	-1	F <sup>-</sup>
Chloride	-1	Cl <sup>-</sup>
Bromide	-1	Br <sup>-</sup>
Iodide	-1	I <sup>-</sup>
Nitride	-3	N <sup>3-</sup>
Oxide	-2	O <sup>2-</sup>
Sulfide	-2	S <sup>2-</sup>
Selenide	-2	Se <sup>2-</sup>
Telluride	-2	Te <sup>2-</sup>
Carbonate	-2	CO <sub>3</sub> <sup>2-</sup>
Bicarbonate	-1	HCO <sub>3</sub> <sup>-</sup>
Phosphate	-3	PO <sub>4</sub> <sup>3-</sup>
Biphosphate	-2	H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>
Hydrogen phosphate	-2	HPO <sub>4</sub> <sup>2-</sup>
Hydrogen sulfide	-1	HS <sup>-</sup>
Sulfate	-2	SO <sub>4</sub> <sup>2-</sup>
Bisulfate	-1	HSO <sub>4</sub> <sup>-</sup>
Perchlorate	-1	ClO <sub>4</sub> <sup>-</sup>
Chlorate	-1	ClO <sub>3</sub> <sup>-</sup>
Sulfate	-2	SO <sub>4</sub> <sup>2-</sup>
Sulfite	-2	SO <sub>3</sub> <sup>2-</sup>
Hydrogen sulfite	-1	HSO <sub>3</sub> <sup>-</sup>
Hydroxide	-1	OH <sup>-</sup>
Hydroxide	-1	OH <sup>-</sup>
Hydroxide	-1	OH <sup>-</sup>

**How to Write Molecular Formulas**

1. Write the symbols for the two elements.

2. Write the number of atoms of each element.

3. Write the formula with the subscripts.

4. Simplify the formula if possible.

Ion	Charge	Symbol
Hydrogen	+1	H <sup>+</sup>
Lithium	+1	Li <sup>+</sup>
Sodium	+1	Na <sup>+</sup>
Potassium	+1	K <sup>+</sup>
Ammonium	+1	NH <sub>4</sub> <sup>+</sup>
Calcium	+2	Ca <sup>2+</sup>
Magnesium	+2	Mg <sup>2+</sup>
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Strontium	+2	Sr <sup>2+</sup>
Lead	+2	Pb <sup>2+</sup>
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Fluoride	-1	F <sup>-</sup>
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Nitride	-3	N <sup>3-</sup>
Oxide	-2	O <sup>2-</sup>
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Carbonate	-2	CO <sub>3</sub> <sup>2-</sup>
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# Writing Ionic Formulas Study Guide

**Karen Timberlakes**



## Writing Ionic Formulas Study Guide:

**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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## **Table of Contents Writing Ionic Formulas Study Guide**

1. Understanding the eBook Writing Ionic Formulas Study Guide
  - The Rise of Digital Reading Writing Ionic Formulas Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Writing Ionic Formulas Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Writing Ionic Formulas Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Writing Ionic Formulas Study Guide
  - Personalized Recommendations
  - Writing Ionic Formulas Study Guide User Reviews and Ratings
  - Writing Ionic Formulas Study Guide and Bestseller Lists
5. Accessing Writing Ionic Formulas Study Guide Free and Paid eBooks
  - Writing Ionic Formulas Study Guide Public Domain eBooks
  - Writing Ionic Formulas Study Guide eBook Subscription Services

- Writing Ionic Formulas Study Guide Budget-Friendly Options
- 6. Navigating Writing Ionic Formulas Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Writing Ionic Formulas Study Guide Compatibility with Devices
  - Writing Ionic Formulas Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Writing Ionic Formulas Study Guide
  - Highlighting and Note-Taking Writing Ionic Formulas Study Guide
  - Interactive Elements Writing Ionic Formulas Study Guide
- 8. Staying Engaged with Writing Ionic Formulas Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Writing Ionic Formulas Study Guide
- 9. Balancing eBooks and Physical Books Writing Ionic Formulas Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Writing Ionic Formulas Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Writing Ionic Formulas Study Guide
  - Setting Reading Goals Writing Ionic Formulas Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Writing Ionic Formulas Study Guide
  - Fact-Checking eBook Content of Writing Ionic Formulas Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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