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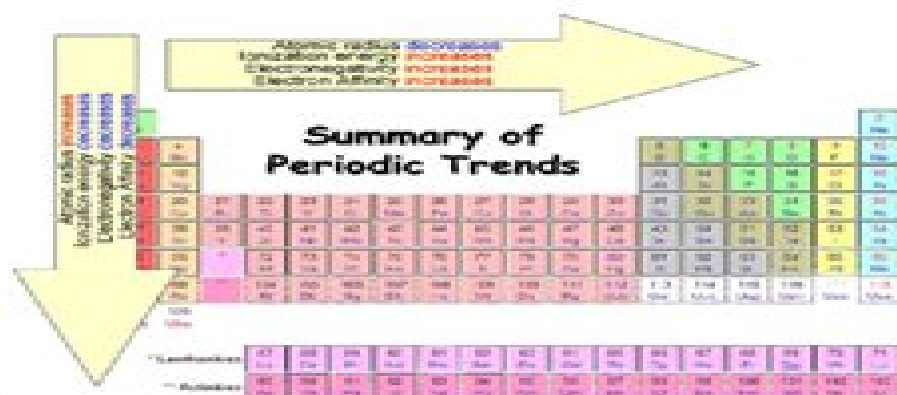
1. Polarizability

↑ SA/ length of , ↑ polarizability

2. IMF

↑ boiling/melting point, ↑ IMF

↑ vapor pressure, ↓ IMF b/c: ↑ mol solute, ↓ pressure & ↓ pressure, ↑ IMF



Trends:

1. Relate strength of intermolecular forces to state of matter

- The state of a substance depends on the balance between the *kinetic energy* of the individual particles (molecules or atoms) and the *intermolecular forces*. The kinetic energy keeps the molecules apart and moving around, and is a function of the temperature of the substance. The intermolecular forces are attractive forces that try to draw the particles together (Figure
- As discussed previously, gasses are very sensitive to temperatures and pressure. However, these also affect liquids and solids too. Heating and cooling can change the *kinetic energy* of the particles in a substance, and so, we can change the physical state of a substance by heating or cooling it. Increasing the pressure on a substance forces the molecules closer together, which *increases* the strength of intermolecular forces.

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Table of Contents Unit 6 Section D Content Guide Chemistry

1. Understanding the eBook Unit 6 Section D Content Guide Chemistry
 - The Rise of Digital Reading Unit 6 Section D Content Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 6 Section D Content Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 6 Section D Content Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 6 Section D Content Guide Chemistry

- Personalized Recommendations
- Unit 6 Section D Content Guide Chemistry User Reviews and Ratings
- Unit 6 Section D Content Guide Chemistry and Bestseller Lists
- 5. Accessing Unit 6 Section D Content Guide Chemistry Free and Paid eBooks
 - Unit 6 Section D Content Guide Chemistry Public Domain eBooks
 - Unit 6 Section D Content Guide Chemistry eBook Subscription Services
 - Unit 6 Section D Content Guide Chemistry Budget-Friendly Options
- 6. Navigating Unit 6 Section D Content Guide Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 6 Section D Content Guide Chemistry Compatibility with Devices
 - Unit 6 Section D Content Guide Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 6 Section D Content Guide Chemistry
 - Highlighting and Note-Taking Unit 6 Section D Content Guide Chemistry
 - Interactive Elements Unit 6 Section D Content Guide Chemistry
- 8. Staying Engaged with Unit 6 Section D Content Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 6 Section D Content Guide Chemistry
- 9. Balancing eBooks and Physical Books Unit 6 Section D Content Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 6 Section D Content Guide Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 6 Section D Content Guide Chemistry
 - Setting Reading Goals Unit 6 Section D Content Guide Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 6 Section D Content Guide Chemistry

- Fact-Checking eBook Content of Unit 6 Section D Content Guide Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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