

Topic 7.7: Calculating Equilibrium Concentrations

- The concentrations or partial pressures of species at equilibrium can be predicted given the balanced equation, initial concentrations, and the appropriate K .

EXAMPLE

Consider the reaction represented by the equation:



- Trial #1:

6.00 M $\text{Fe}^{3+}(\text{aq})$ and 10.0 M $\text{SCN}^-(\text{aq})$ are mixed at a certain temperature and at equilibrium the concentration of $\text{FeSCN}^{2+}(\text{aq})$ is 4.00 M.

What is the value for the equilibrium constant for this reaction?

	Fe^{3+}	SCN^-	FeSCN^{2+}
Initial			
Change			
Equilibrium			

- Set up the ICE table with the given values.
- Fill in the changes.
- Calculate the equilibrium values.
- Plug Eq values into the K expression.

Unit 7 Chemistry Unit Guide Key

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Estridge,Ruth Reynolds,2000-05 *Cooperative Chemistry Lab Manual* Cooper,2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations *The Cumulative Book Index* ,1913 **National Laundry and Cleaning Journal** ,1920 The British National Bibliography Arthur James Wells,1996 **Books in Print** ,1982 Resources in Education ,1995
Journal of Chemical Education ,1928 *Books and Pamphlets, Including Serials and Contributions to Periodicals* Library of Congress. Copyright Office,1974 **Basic Medical Laboratory Techniques** Norma J. Walters,1991 Whitaker's Five-year Cumulative Book List ,1968 *Catalogue of Central Wesleyan College, Warrenton, Mo.* ... Central Wesleyan College (Warrenton, Mo.),1917 **Year-book** University of Southern California,1918 **Forthcoming Books** Rose Arny,2003

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