

## Unit 7 Review Sheet (Stoichiometry)

1.  $2\text{KClO}_3 \rightarrow 3\text{O}_2 + 2\text{KCl}$   

$$\frac{80.5\text{g O}_2}{32.0\text{g O}_2} \times \frac{1\text{mol O}_2}{3\text{mol O}_2} \times \frac{2\text{mol KCl}}{1\text{mol KCl}} \times \frac{74.6\text{g KCl}}{1\text{mol KCl}} = \boxed{125\text{g KCl}}$$
- 2.)  $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$   

$$\frac{1.44\text{g Na}}{23.0\text{g Na}} \times \frac{1\text{mol Na}}{4\text{mol Na}} \times \frac{2\text{mol Na}_2\text{O}}{1\text{mol Na}_2\text{O}} \times \frac{62.0\text{g Na}_2\text{O}}{1\text{mol Na}_2\text{O}} = \boxed{1.94\text{g Na}_2\text{O}}$$
- 3.)  $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$   

$$\frac{2.55\text{g Zn}}{65.4\text{g Zn}} \times \frac{1\text{mol Zn}}{1\text{mol Zn}} \times \frac{1\text{mol H}_2}{1\text{mol H}_2} \times \frac{22.4\text{L H}_2}{1\text{mol H}_2} = \boxed{0.873\text{L H}_2}$$
- 4.)  $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$   

$$\frac{57.1\text{g Fe}_2\text{O}_3}{159.6\text{g Fe}_2\text{O}_3} \times \frac{1\text{mol Fe}_2\text{O}_3}{1\text{mol Fe}_2\text{O}_3} \times \frac{3\text{mol CO}_2}{1\text{mol CO}_2} \times \frac{22.4\text{L CO}_2}{1\text{mol CO}_2} = \boxed{24.0\text{L CO}_2}$$
- 5)  $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$   

$$\frac{33.0\text{g Na}}{23.0\text{g Na}} \times \frac{1\text{mol Na}}{2\text{mol Na}} \times \frac{2\text{mol NaCl}}{1\text{mol NaCl}} \times \frac{58.5\text{g NaCl}}{1\text{mol NaCl}} = 83.9\text{g NaCl}$$
  

$$\frac{11.2\text{L Cl}_2}{22.4\text{L Cl}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol Cl}_2} \times \frac{2\text{mol NaCl}}{1\text{mol NaCl}} \times \frac{58.5\text{g NaCl}}{1\text{mol NaCl}} = \boxed{58.5\text{g NaCl}}$$

# Unit 7 Stoichiometry Review

**Robert A. Huggins**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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