

During this practical session, find the molar ratio between silver nitrate and

AgNO<sub>3</sub> and AgCl

1. Balance the equation
2. If you have 1.25 grams silver and 1.50 grams of Ag<sub>2</sub>, what is the limiting reagent?
3. After finding the answer to Q2, determine mass of AgCl that should be produced (theoretical yield)
4. Convert 1.50 grams silver to moles of Ag
5. If this number is now written vertically is 1.50 grams of AgCl actually the 1.50 grams?

# Unit 7 Stoichiometry Study Guide

**Darrell D. Ebbing**



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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Unit 7 Stoichiometry Study Guide** a literary masterpiece penned with a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those that partake in its reading experience.

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