

- $1.0 \times 10^{-14}$  is the equilibrium constant for the autoionization of water ( $K_w$ )
- Equilibrium occurs in all solutions (acid, base, neutral)
  - All aqueous solutions have a detectable concentration of both hydronium and hydroxide ions
  - The product of all of these ion concentrations is always  $K_w$
  - $K_w = H_3O^{+} \cdot OH^{-} = 1.0 \times 10^{-14}$
- If concentration of hydroxide ion is known, to determine the pOH:
  - $pOH = 14 - pH$
- If pOH is known, to determine pH:
  - $pH = 14 - pOH$
- If the hydronium concentration is known, to determine the hydroxide ion concentration :
  - $(OH^{-}) = 1.0 \times 10^{-14} / (H_3O^{+})$
- Calculate pH from the hydronium concentration:
  - $pH = -\log (H_3O^{+})$
- If pH is known, to determine the hydronium concentration:
  - $(H_3O^{+}) = 10^{-pH}$

### Buffer Solutions

- A weak acid or base mixed with its conjugate
- It resists a change in pH when H<sup>+</sup> or OH<sup>-</sup> is added
  - when H<sup>+</sup> or OH<sup>-</sup> is added to the buffer:
    - Decreases relative buffer component ratio
    - Works out the stoichiometry to find the new concentrations of buffer components
    - Use equilibrium constant to determine the new pH
- Buffers in acidic solutions
  - Weak acid + Conjugate Base
  - Weak Base + Conjugate Acid
  - Excess Weak Acid + Strong Base
  - Excess Weak Base + Strong Acid

# Unit 8 Chemistry Study Guide Anwsers

**L Reisser**

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