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Louis Nashelsky



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 theory and polarity name date formula number of lewis dot electrons formula molecular electron geometry drawing of
 structure hybridization geometry hinh trigonal 8 tetrahedral sp^3 pyramidal nh_3 h fb f

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 lewis structures of these molecules by adding multiple bonds and lone pairs do not add any more atoms a the amino acid
 serine b urea c pyruvic acid d uracil e carbonic acid

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 figure 1 a lewis structure figure 1a is meant to represent chemical bonding and the distribution of electrons

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octet rule where each atom has a full valence shell with eight electrons

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web the pcl 3 lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for nh 3 chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom when drawing the structure of an ion be sure to add subtract electrons to account for the charge step 2 connect the atoms to each other with single bonds to form a skeleton structure

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web this exercise required us to take the formulas in the data table and calculate the number of valence electrons using the periodic table draw its lewis structure on paper build the molecular model and record a photo list the number of bonds and lone pairs surrounding the central atom and lastly identify the structure geometry

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