

Chapter 8 - Solution Stoichiometry - Multiple Choice Questions

- Solubility is a term:
 - that always means that more than 1g of solute will dissolve in 100 mL of water.
 - that can be used for all solutes because they all dissolve to some extent.
 - that generally means that at least 1g of solute will dissolve in 100 mL of water.
 - that is used in the same way by all chemists.
- Which statement is true?
 - When an atom loses an electron the cation formed is smaller and when an atom gains an electron the anion formed is larger.
 - When an atom gains or loses an electron the ion formed is smaller.
 - When an atom gains or loses an electron the ion formed is larger.
 - When an atom loses an electron the cation formed is larger and when an atom gains an electron the anion formed is smaller.
- Two important factors that affect the solubility of an ionic compound are:
 - ion charge and agitation.
 - ion size and agitation.
 - ion size and ion charge.
 - ion size and the particle size of the compound.
- Which statement is true?
 - Compounds with small ions tend to be more soluble than compound with large ions.
 - Compounds with large ions tend to be less soluble than compounds with small ions.
 - Compounds with small ionic charges tend to be less soluble than compounds with large ionic charges.
 - Compounds with large ions tend to be more soluble than compounds with small ions.
- Which anion when combined with Mg^{2+} will probably give the most soluble ionic compound?
 - F^{-}
 - Cl^{-}
 - Br^{-}
 - I^{-}
- Which cation when combined with oxide ion (O^{2-}) is most likely to give a soluble compound?
 - Ca^{2+}
 - K^{+}
 - Al^{3+}
 - Co^{2+}
- According to the general solubility guidelines, which compound is not soluble?
 - $CaCO_3$
 - $LiBr$
 - $CaCl_2$
 - $Mg(NO_3)_2$
- According to the general solubility guidelines, which compound is soluble?
 - $BeSO_4$
 - Na_2SO_4
 - Ag_2SO_4
 - $CaSO_4$
- In a double displacement reaction:
 - two ionic compounds react by exchanging ions.
 - a precipitate may be formed.
 - a gas may be produced.
 - all of the above.
- Which combination of solutions will result in the formation of a precipitate?
 - $Ba(NO_3)_2(aq) + Na_2CO_3(aq)$
 - $KNO_3(aq) + Na_2CO_3(aq)$
 - $Ca(NO_3)_2(aq) + NaCl(aq)$
 - $NH_4NO_3(aq) + NaCl(aq)$
- A metal carbonate reacts with an acid to produce a gas. Which gas is formed?
 - $N_2(g)$
 - $CO_2(g)$
 - $CO(g)$
 - $H_2O(g)$
- Which of the following is one of the products of a reaction between an acid and a base?
 - Water
 - Hydronium ion
 - Hydroxide ion
 - Hydrogen gas
- Identify the spectator ions in the following total ionic equation:
 $Na^+(aq) + OH^-(aq) + Ag^+(aq) + NO_3^-(aq) \rightarrow Na^+(aq) + AgOH(s) + NO_3^-(aq)$
 - Na^{+} and OH^{-}
 - Ag^{+} and $[NO_3]^{-}$
 - OH^{-} and Ag^{+}
 - Na^{+} and $[NO_3]^{-}$
- The following equation is an example of what type of equation?
 $2Ag^+(aq) + S^{2-}(aq) \rightarrow Ag_2S(s)$
 - Molecular equation
 - Total ionic equation
 - Net ionic equation
 - Skeleton equation
- In writing the total ionic equation for the reaction between sodium phosphate and copper(II) chloride, which compound is not written as it's dissociated ions?
 - Sodium chloride
 - Copper(II) phosphate
 - Copper(II) chloride
 - Sodium phosphate
- The net ionic equation for the reaction that occurs from mixing a solution of barium chloride with a solution of potassium sulfate is:
 - $K^+(aq) + Cl^-(aq) \rightarrow KCl(aq)$
 - $BaCl_2(aq) \rightarrow Ba^{2+}(aq) + 2Cl^-(aq)$
 - $Ba^{2+}(aq) + SO_4^{2-}(aq) \rightarrow BaSO_4(s)$
 - $BaSO_4(s) \rightarrow Ba^{2+}(aq) + SO_4^{2-}(aq)$
- What is qualitative analysis?
 - It determines how much of a compound, element or ion is in an unknown sample.
 - It identifies compounds, elements or ions in an unknown solution.
 - It measures concentrations of ions.
 - None of the above.
- What is the dissociation equation for an aqueous solution of magnesium chloride?
 - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl_2(aq)$
 - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl^-(aq)$
 - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + Cl_2(aq)$
 - $MgCl_2(aq) \rightarrow Mg^{2+}(aq) + 2Cl^-(aq)$
- What is the concentration of nitrate ions in a 0.15 mol/L solution of $Fe(NO_3)_3$?
 - 0.45 mol/L
 - 0.30 mol/L
 - 0.15 mol/L
 - There is not enough information to answer the question.
- What is the concentration of a solution of aluminum sulfate ($Al_2(SO_4)_3$) if the concentration of sulfate ions is 0.60 mol/L?
 - 1.2 mol/L
 - 0.60 mol/L
 - 0.40 mol/L
 - 0.20 mol/L
- Equal volumes of 0.20 mol/L $NaCl$ and 0.30 mol/L KCl are mixed together. What is the concentration of chloride ions in the mixture?
 - 0.20 mol/L
 - 0.25 mol/L
 - 0.30 mol/L
 - You need to know the volumes of the original solutions to answer the question.
- What volume of a 0.500 mol/L solution of Na_2CO_3 contains enough carbonate ions to completely precipitate 0.0550 mol of Ba^{2+} ions?
 - 300 mL
 - 90.0 mL
 - 30.0 mL
 - 10.0 mL
- What volume of a 1.50 mol/L solution of $Ba(OH)_2$ contains enough hydroxide ions to completely precipitate the iron(III) ions found in 60.0 mL of a 0.500 mol/L solution of $FeCl_3$?
 - 180 mL
 - 60.0 mL
 - 20.0 mL
 - 12.5 mL
- How many moles of precipitate $BaSO_4$ are formed when 50.0 mL of a 0.500 mol/L solution of $BaCl_2$ is mixed with 35.0 mL of a 0.500 mol/L solution of Na_2SO_4 ?
 - 0.0175 mol
 - 0.0250 mol
 - 0.0050 mol
 - 0.0800 mol
- How many moles of $Mg(OH)_2$ will be precipitated when 75.0 mL of 0.400 mol/L $MgCl_2$ is mixed with 75.0 mL of 0.400 mol/L $NaOH$?
 - 0.0750 mol
 - 0.0600 mol
 - 0.0000 mol
 - 0.0050 mol
- In the above question, what mass of precipitate is formed?
 - 0.875 g
 - 1.75 g
 - 3.50 g
 - 4.37 g

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