

# Preparation of Solution

## Molar Solutions

What do you know?

Mass of Solute (g)

Final Volume of Solution (L)

$$\text{Mass of Solute (g)} = \frac{1}{\text{Molar Mass } \left(\frac{\text{g}}{\text{mole}}\right)} \times \frac{1}{\text{Final Volume of solution (L)}} = \text{Molarity (M or } \frac{\text{mole}}{\text{L}})$$

What do you want to know?

Molar Mass of Solute (M)

Final Volume of Solution (L)

$$\text{Molarity } \left(\frac{\text{mole}}{\text{L}}\right) = \frac{\text{Mass of Solute (g)}}{\text{Molar Mass } \left(\frac{\text{g}}{\text{mole}}\right)} \times \frac{1}{\text{Final Volume of solution (L)}} = \text{Mass of Solute (g)}$$

What do you want to know?

Mass of Solute (g)

Molar Mass of Solute (M)

$$\text{Mass of Solute (g)} = \frac{1}{\text{Molar Mass } \left(\frac{\text{g}}{\text{mole}}\right)} \times \frac{1}{\text{Molarity (M or } \frac{\text{mole}}{\text{L}})} = \text{Final Volume of Solution (L)}$$

What do you know?

Mass of Solute (g)

Molar Mass of Solute (M)

What do you want to know?

Final Volume of Solution (L)

## Percent (%) Solutions

Mass/Mass Percent Concentration Solution - All percentages should be in decimal form

What do you know?

Mass of Solute (g)

Mass of Solvent (g)

$$\frac{\text{Mass of Solute (g)}}{\text{Mass of Solute (g)} + \text{Mass of Solvent (g)}} = \text{Percent Concentration of Solution (\%)}$$

What do you want to know?

Percent Concentration of Solution (%)

What do you know?

Mass of Solute (g)

Percent Concentration of Solution (%)

$$\frac{\text{Mass of Solute (g)}}{\text{Percent Concentration of Solution (\%)}} = \text{Mass of Solvent (g)} + \text{Mass of Solute (g)}$$

What do you want to know?

Mass of Solvent (g)

What do you know?

Mass of Solvent (g)

Percent Concentration of Solution (%)

$$\frac{\text{Percent Concentration of Solution (\%)}}{1 - \text{Percent Concentration of Solution (\%)}} = \text{Mass of Solute (g)}$$

What do you want to know?

Mass of Solute (g)

Mass/Volume Percent Solutions - All percentages should be in decimal form

What do you know?

Mass of Solute (g)

Final Volume of Solution (mL)

$$\frac{\text{Mass of Solute (g)}}{\text{Final Volume of Solution (mL)}} = \text{Percent Concentration of Solution (\%)}$$

What do you want to know?

Percent Concentration of Solution (%)

What do you know?

Mass of Solute (g)

Percent Concentration of Solution (%)

$$\frac{\text{Mass of Solute (g)}}{\text{Percent Concentration of Solution (\%)}} = \text{Final Volume of Solution (mL)}$$

What do you want to know?

Final Volume of Solution (mL)

What do you know?

Final Volume of Solution (mL)

Percent Concentration of Solution (%)

$$\text{Final Volume of Solution (mL)} = \frac{\text{Mass of Solute (g)}}{\text{Percent Concentration of Solution (\%)}}$$

What do you want to know?

Mass of Solute (g)

## Dilutions

Molar Solution:

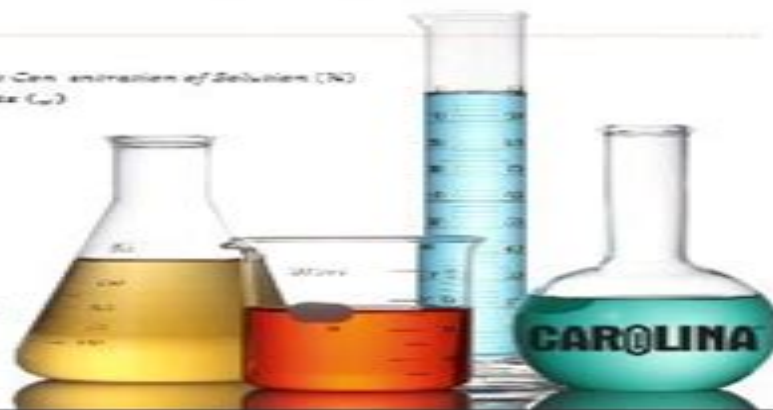
Molarity of Initial Solution  $\left(\frac{\text{mole}}{\text{L}}\right) \times \text{Volume of Initial Solution (L)} =$

Molarity of Final Solution  $\left(\frac{\text{mole}}{\text{L}}\right) \times \text{Volume of Final Solution (L)}$

Percent Concentration Solutions

Percent Concentration of Initial Solution  $\left(\frac{\text{g}}{\text{mL}}\right) \times \text{Volume of Initial Solution (L)} =$

Percent Concentration of Final Solution  $\left(\frac{\text{g}}{\text{mL}}\right) \times \text{Volume of Final Solution (L)}$



# Solution Preparation Manual Chemistry

**Paul A. Bartlett, Judith G. Koch**



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