

SALT ANALYSIS

TEST FOR ANIONS

EXPERIMENT	OBSERVATION	INFERENCE
1. Test for CO_3^{2-} → A little of the salt is treated with dil. HCl.	Brisk effervescence with the evolution of colourless odourless gas	May be CO_3^{2-} $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
Confirmatory test :- 2. Pass the above gas through lime water.	Lime water turns milky	Presence of CO_3^{2-} is confirmed $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$
Test for Halides 3. A little of the salt is treated with 1 or 2 drops of conc. H_2SO_4 .	a) Colourless pungent smelling gas evolves, dense white fumes when a glass rod is dipped in NH_4OH is shown on the mouth of the test tube b) Reddish brown vapours with pungent smell c) Deep violet vapours with pungent smell	a) May be Cl^- $\text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{NaHSO}_4 + \text{HCl}$ b) Maybe Br^- $\text{NaBr} + \text{H}_2\text{SO}_4 \rightarrow \text{NaHSO}_4 + \text{HBr}$ $\text{H}_2\text{SO}_4 + 2\text{HBr} \rightarrow \text{SO}_2 + \text{Br}_2 + 2\text{H}_2\text{O}$ c) May be I^- $\text{KI} + \text{H}_2\text{SO}_4 \rightarrow \text{KHSO}_4 + \text{HI}$ $\text{H}_2\text{SO}_4 + 2\text{HI} \rightarrow \text{SO}_2 + \text{I}_2 + 2\text{H}_2\text{O}$
Confirmatory test of halides 4. A little of the salt solution is acidified with dil. HNO_3 , boil, cool and add AgNO_3 sol.	a) Curdy white ppt. soluble in NH_4OH . b) Pale yellow ppt. sparingly soluble in NH_4OH c) Yellow ppt. insoluble in NH_4OH	a) presence of Cl^- is confirmed $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}$ (white ppt.) b) presence of Br^- is confirmed $\text{KBr} + \text{AgNO}_3 \rightarrow \text{KNO}_3 + \text{AgBr}$ (pale yellow ppt.) c) presence of I^- is confirmed $\text{KI} + \text{AgNO}_3 \rightarrow \text{KNO}_3 + \text{AgI}$ (yellow ppt.)

Unknown Salt Identification

David Collins (Ph. D.), Matthew Johll

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