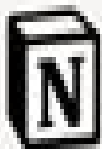


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# Morning Routine Guide



@elhnail



## My Morning Routine Guide

Morning Routine Kit



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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 63.9 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{100 \text{ g}}{160.0 \text{ g}} \times 100 = 62.5\%$

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