

Regents Chemistry Final Exam Review

Topic 4- Properties of Solutions

Define the following terms:

**Saturated:** An amount of solute dissolved is equal to the maximum the solvent can hold

**Unsaturated:** Not enough solute has been dissolved

**Supersaturated:** Extra solute has been dissolved

**Homogeneous:** The same throughout (Ex: Salt water)

**Heterogeneous:** Different throughout (Ex: Oil and water)

3. Explain how the following factors affect solubility:

a) Temperature

b) Pressure

c) "Like dissolves like"

4. Using Table G, answer the following questions:

a) What substances from Table G are gases?

b) How much  $\text{KNO}_3$  is required to produce a saturated solution at  $50^\circ\text{C}$  in 200 g of water?

c) If 50 grams of  $\text{HCl}$  is added to 300 g of water at  $20^\circ\text{C}$ , is the solution unsaturated, saturated, or supersaturated?

d) How many grams of Sodium Nitrate is required to produce a saturated solution at  $50^\circ\text{C}$  in 200 g of water?

e)  $\text{KNO}_3$  is added to 100 g of water to produce a saturated solution at  $40^\circ\text{C}$ , if the solution is cooled to  $20^\circ\text{C}$ , how many grams of  $\text{KNO}_3$  will precipitate out?

## Topic 7 Properties Of Solutions Review Questions

**James E. Brady, Gerard E. Humiston**

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