

Mole Review Problems

Solve the following problems and SHOW YOUR WORK

1. How many moles of ATOMS are there in 1.20×10^{25} methane (CH_4) molecules?

$$1.20 \times 10^{25} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) \left(\frac{5 \text{ atoms}}{1 \text{ molecule}} \right) = 99.7 \text{ moles of atoms}$$

2. Calculate the molarity of a solution formed by dissolving 32.5 g of NaCl in sufficient water to make 2.5 L of solution.

$$\text{FW NaCl} = 23.0 + 35.5 = 58.5 \text{ g/mol}$$

$$32.5 \text{ g} \div 58.5 \text{ g/mol} = 0.556 \text{ mol}$$

$$M = \frac{n}{V} = \frac{0.556 \text{ mol}}{2.5 \text{ L}} = 0.22 \text{ M}$$

3. How many ATOMS are contained in 32.0 grams of gaseous oxygen?

$$\text{FW O}_2 = (16.0) = 32.0 \text{ g/mol}$$

$$32.0 \text{ g} \div 32.0 \text{ g/mol} = 1.00 \text{ mol O}_2$$

$$1.00 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \times \frac{2 \text{ atoms}}{1 \text{ O}_2 \text{ molecule}} = 1.20 \times 10^{24} \text{ atoms}$$

4. 0.500 mole of an element has a mass of 31.75 grams. Identify the element.

$$\text{FW} = \frac{31.75 \text{ g}}{0.500 \text{ mol}} = 63.5 \text{ g/mol} \quad \text{The element is copper.}$$

5. Calculate the number of moles of neon gas in 54 litres at STP.

$$54 \text{ L} \div 22.4 \text{ L/mol} = 2.4 \text{ mol}$$

6. A sample of carbon reacts with a 9.2 gram sample of hydrogen to form a gaseous compound that has a mass of 120.0 grams. What is the molecular formula of this hydrocarbon, which has a molar mass of 26.0 grams?

$$\text{H: } 9.2 \text{ g} \div 1.09 \text{ g/mol} = 9.2 \text{ moles H} \quad \text{Mole Ratio: C:H} \approx 1:1$$

$$\text{C: } (120.0 \text{ g} - 9.2 \text{ g}) \div 12.09 \text{ g/mol} = 9.2 \text{ moles C} \quad \text{Molecular mass} \approx 26.0 \text{ g/mol}$$

Empirical formula is $\text{CH} \rightarrow \text{molar mass} \approx 12.0 + 1.0 = 13.0 \text{ g/mol}$ so formula is C_2H_4

7. What is the mass of 2.11×10^{24} molecules of chlorine gas?

$$2.11 \times 10^{24} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = 3.50 \text{ moles Cl}_2$$

$$\text{FW Cl}_2 = 2(35.5) = 71.0 \text{ g/mol}$$

$$3.50 \text{ moles} \times 71.0 \text{ g/mol} = 249 \text{ g}$$

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Kansas State Agricultural College

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Table of Contents Skill Practice 51 Answers Chemistry

1. Understanding the eBook Skill Practice 51 Answers Chemistry
 - The Rise of Digital Reading Skill Practice 51 Answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 51 Answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 51 Answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 51 Answers Chemistry
 - Personalized Recommendations
 - Skill Practice 51 Answers Chemistry User Reviews and Ratings
 - Skill Practice 51 Answers Chemistry and Bestseller Lists
5. Accessing Skill Practice 51 Answers Chemistry Free and Paid eBooks
 - Skill Practice 51 Answers Chemistry Public Domain eBooks
 - Skill Practice 51 Answers Chemistry eBook Subscription Services
 - Skill Practice 51 Answers Chemistry Budget-Friendly Options

6. Navigating Skill Practice 51 Answers Chemistry eBook Formats
 - ePUB, PDF, MOBI, and More
 - Skill Practice 51 Answers Chemistry Compatibility with Devices
 - Skill Practice 51 Answers Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 51 Answers Chemistry
 - Highlighting and Note-Taking Skill Practice 51 Answers Chemistry
 - Interactive Elements Skill Practice 51 Answers Chemistry
8. Staying Engaged with Skill Practice 51 Answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 51 Answers Chemistry
9. Balancing eBooks and Physical Books Skill Practice 51 Answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 51 Answers Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Skill Practice 51 Answers Chemistry
 - Setting Reading Goals Skill Practice 51 Answers Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Skill Practice 51 Answers Chemistry
 - Fact-Checking eBook Content of Skill Practice 51 Answers Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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