

NAME: _____

Molarity Practice

Date: _____

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1. Calculate the molarity of the following solutions:

a) 45.0 g of NaOH in 2.00 L of solution.b) 10.0 g of CH_3COOH in 25 mL of solution.c) 175.0 g of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ in 500 mL of solution.

2. What is the molarity of a solution that contains 25.0 g of each of the following?

a) 25.0 g of NaOH b) 25.0 g of Na_2CO_3

3. What is the molarity of a solution that contains 10.0 g of each of the following in 250 mL of solution?

4. Which of the following contains the highest concentration? Show using calculations.

a) 45.0 g of NaOH in 2.00 L of solutionb) 25.0 g of NaOH in 500 mL

5. Make a more concentrated solution of sodium hydroxide (500 mL) of solution by pouring a solution that has a concentration of 1.0 M. What you need to make this solution? Show the molarity and the mass, in that the solution is already prepared.

Skill Practice 53 Answers Chemistry

California, Carter Pitkin Pomeroy



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