

"Mastering Chemical Equilibrium: A Comprehensive Study Guide"

Chemical Equilibrium Study Guide:

I. Introduction:

- **Definition:** State where the rate of the forward and reverse reactions are equal.
- **Significance:** Understanding chemical equilibrium helps in explaining the behavior of reactions under different conditions.

II. Factors Affecting Chemical Equilibrium:

- Concentration of reactants and products
- Temperature
- Pressure
- Catalyst

III. Equilibrium Constant (K_c):

- **Definition:** Measure of the relative concentrations of reactants and products at equilibrium.
- **Calculation:** $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
- **Significance:** K_c can be used to predict the direction of the reaction under different conditions.

IV. Le Chatelier's Principle:

- **Definition:** States that a system at equilibrium will counteract any disturbance to maintain the balance.
- **Application:** Can be used to predict the effect of changes in concentration, temperature, pressure, and presence of catalyst on the equilibrium.

A. Concentration and Equilibrium:

- Increasing the concentration of reactants or decreasing the concentration of products will shift the reaction towards the side with fewer moles.
- Decreasing the concentration of reactants or increasing the concentration of products will shift the reaction towards the side with more moles.

B. Temperature and Equilibrium:

- Endothermic reactions: Absorption of heat shifts the equilibrium to the right with increasing temperature.
- Exothermic reactions: Release of heat shifts the equilibrium to the left with increasing temperature.

C. Pressure and Equilibrium:

- Gases: Increasing the pressure shifts the reaction towards the side with fewer moles of gas.

Study Guide Chemical Equilibrium

Steven S. Zumdahl, Paul B. Kelter



Study Guide Chemical Equilibrium:

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