

Ex 1. For which of the following reactions would the ΔH° and the equilibrium constant K_{eq} be

1) $\Delta H^\circ_{rxn} < 0$ & $K_{eq} > 1$ \rightarrow exothermic

2) $\Delta H^\circ_{rxn} < 0$ & $K_{eq} < 1$ \rightarrow endothermic

3) $\Delta H^\circ_{rxn} > 0$ & $K_{eq} > 1$ \rightarrow exothermic

4) $\Delta H^\circ_{rxn} > 0$ & $K_{eq} < 1$ \rightarrow endothermic

Ex 2. The ΔH° for the reaction between acetylene and hydrogen is:

The following thermodynamic information:

$\text{H}_2\text{C}=\text{CH}_2(g) + \text{H}_2(g) \rightarrow \text{CH}_3\text{CH}_3(g) \quad \Delta H^\circ = -137 \text{ kJ}$

$\text{H}_2\text{C}=\text{CH}_2(g) \rightarrow \text{H}_2\text{C}=\text{CH}(g) + \text{H}(g) \quad \Delta H^\circ = +226 \text{ kJ}$

Calculate the enthalpy of formation of $\text{H}_2\text{C}=\text{CH}_2$ in the form:

$\text{H}_2\text{C}=\text{CH}_2(g) + \text{H}_2(g) \rightarrow \text{H}_2\text{C}=\text{CH}(g) + \text{H}(g)$

$\Delta H^\circ_{rxn} = \text{kJ}$ $\Delta H^\circ_{rxn} = \text{kJ}$ $\Delta H^\circ_{rxn} = \text{kJ}$

The answer is: $\Delta H^\circ_{rxn} = \text{kJ}$

Ex 3. Calculate ΔH°_{rxn} for the following reaction in kJ mol^{-1} :

$\text{C}_2\text{H}_4(g) + \text{H}_2(g) \rightarrow \text{C}_2\text{H}_6(g)$

$\Delta H^\circ_{f, \text{C}_2\text{H}_4(g)} = 52.3 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{C}_2\text{H}_6(g)} = -84.7 \text{ kJ mol}^{-1}$

$\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$

Ex 4. Calculate the heat of formation of $\text{H}_2\text{O}(l)$ for the reaction between $\text{H}_2(g)$ and $\text{O}_2(g)$ to form $\text{H}_2\text{O}(l)$.

$\text{H}_2(g) + \text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$

Reaction Energy

$\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$

$\Delta H^\circ_{f, \text{O}_2(g)} = 0 \text{ kJ mol}^{-1}$

$\Delta H^\circ_{f, \text{H}_2\text{O}(l)} = -285.8 \text{ kJ mol}^{-1}$

$\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{O}_2(g)} = 0 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{H}_2\text{O}(l)} = -285.8 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{H}_2\text{O}(l)} = -285.8 \text{ kJ mol}^{-1}$

Ex 5. Calculate the standard heat of formation $\Delta H^\circ_{f, \text{H}_2\text{O}(l)}$ for the following information:

$\text{H}_2\text{O}(l) + \text{H}_2(g) \rightarrow \text{H}_2\text{O}(g) + \text{H}_2(g)$

$\Delta H^\circ_{rxn} = -483.6 \text{ kJ mol}^{-1}$

$\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$

$\Delta H^\circ_{f, \text{H}_2\text{O}(g)} = -241.8 \text{ kJ mol}^{-1}$

$\Delta H^\circ_{f, \text{H}_2(g)} = 0 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{O}_2(g)} = 0 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{H}_2\text{O}(l)} = -285.8 \text{ kJ mol}^{-1}$ $\Delta H^\circ_{f, \text{H}_2\text{O}(l)} = -285.8 \text{ kJ mol}^{-1}$

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Dixie Goss



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