



#### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay

#### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



#### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- Sulfur hexafluoride  $\text{SF}_6$
- Lithium nitride  $\text{Li}_3\text{N}$
- Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- Tin (II) chloride  $\text{SnCl}_2$
- Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- Potassium bromate  $\text{KBrO}_3$
- Hydrosulfuric acid  $\text{H}_2\text{S}(\text{aq})$
- Chloric acid  $\text{HClO}_3$
- Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- $\text{CuSO}_4$  Copper (II) sulfate
- $\text{AlF}_3$  aluminum fluoride
- $\text{HI}$  hydrogen iodide
- $\text{NO}$  nitrogen monoxide
- $\text{H}_2\text{Se}$  hydroseonic acid
- $\text{HNO}_3$  nitric acid
- $\text{HNO}_2$  nitrous acid
- $\text{NaHSO}_4$  sodium bisulfate

#### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

#### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

# Study Guide Answers Chemistry Key

**H. Stephen Stoker,Danny V. White**

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## **Table of Contents Study Guide Answers Chemistry Key**

1. Understanding the eBook Study Guide Answers Chemistry Key
  - The Rise of Digital Reading Study Guide Answers Chemistry Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Answers Chemistry Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Answers Chemistry Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Answers Chemistry Key
  - Personalized Recommendations
  - Study Guide Answers Chemistry Key User Reviews and Ratings
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- Study Guide Answers Chemistry Key Enhanced eBook Features

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- Highlighting and Note-Taking Study Guide Answers Chemistry Key
- Interactive Elements Study Guide Answers Chemistry Key

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- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide Answers Chemistry Key

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- Benefits of a Digital Library
- Creating a Diverse Reading Collection Study Guide Answers Chemistry Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Study Guide Answers Chemistry Key

- Setting Reading Goals Study Guide Answers Chemistry Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Study Guide Answers Chemistry Key

- Fact-Checking eBook Content of Study Guide Answers Chemistry Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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