

# Biochemistry

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## Chapter 1: Overview

1. **Biochemistry has unity because of biological evolution**
2. **Biology uses mainly six elements plus inorganic salts**
  - 65% Oxygen, 18% Carbon, 10% Hydrogen, 3% Nitrogen, 2% Calcium (salt), 1.1% Phosphorus, 0.35% Potassium
  - Small amounts of S, Na, Cl, Mg, Fe, I
  - Water (60-80% of our mass), it participates in reactions and is a limited solvent
3. **Biochemistry consists of five classes of organic compounds**
  1. Carbohydrates
  2. Lipids
  3. Nucleic Acids
  4. Proteins
  5. Small organic molecule, Inorganic molecules & ions
4. **DNA provides an example of how structure determines function**
  - DNA is very polar
5. Important biochemical interactions include electrostatic interactions, hydrogen bonds, and hydrophobic interactions
  - Electrostatic attractions
  - Hydrogen bonds: involve very strong partial charges and stronger than the average polar bond
  - Hydrophobic interactions (London dispersion/ van der Waals): most important because most abundant. A water "poor" area created by hydrophobic interactions
6. **Important types of reactions in biochemistry include redox and acid-base reactions**
  - a. Redox reactions
  - b. Acid-base reactions
    - Bronsted- Lowry:
    - acids = proton donor
    - bases = proton acceptor
    - Moiety: an atom or group of atoms that composes a molecule or part of a molecule. ( $pK_a$  is the acidity or basicity of a moiety)
    - A moiety is protonated when the pH is lower than its  $pK_a$ , and unprotonated when the pH is higher than its  $pK_a$
7. **Modern biochemistry includes discovery-based as well as hypothesis-driven methods**
  - Discovery based: genetic screening (what gene caused the mutation), biological surveys (what living organisms are in this sector of land)

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**Arshad Iqbal**



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## **Table of Contents Study Guide Biochemistry Test**

1. Understanding the eBook Study Guide Biochemistry Test
  - The Rise of Digital Reading Study Guide Biochemistry Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Biochemistry Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Biochemistry Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Biochemistry Test
  - Personalized Recommendations
  - Study Guide Biochemistry Test User Reviews and Ratings
  - Study Guide Biochemistry Test and Bestseller Lists
5. Accessing Study Guide Biochemistry Test Free and Paid eBooks
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- Study Guide Biochemistry Test Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Study Guide Biochemistry Test Compatibility with Devices
  - Study Guide Biochemistry Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide Biochemistry Test
  - Highlighting and Note-Taking Study Guide Biochemistry Test
  - Interactive Elements Study Guide Biochemistry Test
- 8. Staying Engaged with Study Guide Biochemistry Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide Biochemistry Test
- 9. Balancing eBooks and Physical Books Study Guide Biochemistry Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Biochemistry Test
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Biochemistry Test
  - Setting Reading Goals Study Guide Biochemistry Test
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Biochemistry Test
  - Fact-Checking eBook Content of Study Guide Biochemistry Test
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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