

Comprehensive Study Guide: Chemical Bonding, Matter, Energy, and the Periodic Table

Octet Rule in Chemical Bonding

Atoms tend to react in order to achieve a complete octet of eight valence electrons, matching the electron configuration of noble gases for stability. This drive toward a full outer shell underpins the formation of most chemical bonds and explains why atoms interact in specific ways to attain stable electron arrangements.

Energy Changes in Bond Formation and Breaking

The formation and breaking of chemical bonds involve energy exchanges:

- Bond making releases energy, making it an exothermic process. This energy is given off as atoms come together to form stable bonds.
- Bond breaking requires energy input, making it an endothermic process. Energy must be supplied to overcome the attractive forces holding atoms or ions together.

These energy changes are fundamental to understanding chemical reactions, where the net energy released or absorbed determines whether a process is spontaneous.

Types of Chemical Bonds: Ionic, Covalent, and Metallic

Ionic Bonding

Study Guide Chemical Bonding

**John Krenos, Joseph Potenza, Carl
Hoeger**

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