

Name key

## Physical Science Chemical Bonding Study Guide

### A. Covalent & Ionic Compounds/Bonds

1. When an atom loses an electron it becomes a positive ion (cation), and when it gains an electron it becomes a negative ion (anion).
2. An element in the halogen group has 7 valence electrons, and wants to gain 1 more to have 8. This is because of the octet rule.
3. An ionic bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal.
4. In a covalent bond, atoms share electrons, whereas in an ionic bond, atoms transfer electrons.
5. (Honors only) When electrons are shared unequally in a covalent bond it is polar, and when they're shared equally it is nonpolar. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the electronegativities.
6. An atom can have up to 8 valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called noble gases.
7. Atoms can gain, lose, or share valence electrons to become more stable.

### B. Formulas for Ionic compounds - Fill in the chart below

|                        | Magnesium | Fluorine  | Nitrogen  | Lithium   |
|------------------------|-----------|-----------|-----------|-----------|
| # OF VALENCE ELECTRONS | <u>2</u>  | <u>7</u>  | <u>5</u>  | <u>1</u>  |
| OXIDATION #            | <u>+2</u> | <u>-1</u> | <u>-3</u> | <u>+1</u> |

2. What is the **chemical formula** for each of the following ionic compounds?  
(*hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.*)

- a) Potassium and Chlorine - KCl
- b) Calcium and Bromine - CaBr<sub>2</sub>
- c) Magnesium and oxygen - MgO

# Study Guide Chemical Bonds Answer Key

**Jane B. Taylor**



## **Study Guide Chemical Bonds Answer Key:**

*Study Guide for Chemical Principles* Thomas Elliott Taylor, 1979      *Study Guide for Chemical Principles* [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples      *ATI TEAS 7 Study Guide* Calvin Yates, Are you ready to take the first step toward your nursing career but unsure where to start with preparing for the ATI TEAS 7 exam This comprehensive guide is designed to help you not only understand the structure of the exam but to excel in every section Whether you re a first time test taker or someone looking to improve their score this book offers you the tools and strategies necessary to tackle this critical hurdle in your nursing school application process The ATI TEAS 7 exam is an essential part of your nursing journey assessing your proficiency in key areas such as Reading Mathematics Science and English and Language Usage With detailed explanations test taking strategies and hundreds of practice questions this book offers more than just content review It provides a roadmap for effective studying helping you organize your time and focus on the areas that matter most Inside you ll find a deep dive into each subject area The Reading section teaches you how to improve your comprehension skills identify key details and interpret complex passages The Mathematics section covers everything from basic arithmetic to more advanced concepts like algebra and geometry with strategies to solve problems quickly and accurately In the Science section you ll explore human anatomy biology chemistry and scientific reasoning breaking down complex topics into understandable chunks The English and Language Usage section focuses on grammar sentence structure and vocabulary ensuring you can clearly communicate ideas and effectively navigate language based questions This guide doesn t just help you study it shows you how to study With tips for managing time handling test anxiety and utilizing your calculator effectively you ll learn how to approach the exam strategically You ll gain insight into how to break down multiple choice questions identify key terms and ensure that every answer you choose reflects your true understanding To put everything into practice the book includes 200 practice questions with detailed answers and explanations These questions will test your knowledge strengthen your weaknesses and help you build the confidence you need to succeed This guide is your key to unlocking your potential on the ATI TEAS 7 exam and setting the stage for a successful career in nursing With the right tools dedication and preparation you can face the exam with confidence and take one step closer to your dream      *Class 9 Chemistry MCQ (Multiple Choice Questions)* Arshad Iqbal, *The Class 9 Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF 9th Grade Chemistry MCQ PDF Download Quiz Questions Chapter 1 8 Practice Tests with Answer Key Chemistry Questions Bank MCQs Notes* includes revision guide for problem solving with hundreds of solved MCQs *Class 9 Chemistry MCQ with Answers PDF* book covers basic concepts analytical and practical assessment tests *Class 9 Chemistry MCQ PDF* book helps to practice test questions from exam prep notes *The Class 9 Chemistry MCQs with Answers PDF eBook* includes revision guide with verbal quantitative and analytical past papers solved

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