

**CHAPTER 15** STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

| solvent   | liquid | solvent | solvent |
|-----------|--------|---------|---------|
| insoluble | solid  | solid   | solid   |

After reading (1) \_\_\_\_\_ solution \_\_\_\_\_ of solute is dissolved in solvent.

(2) \_\_\_\_\_ solute \_\_\_\_\_ is the (3) \_\_\_\_\_ solvent \_\_\_\_\_ and solute is the

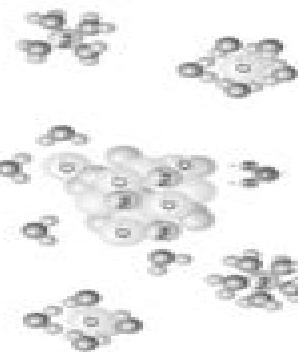
(4) \_\_\_\_\_ solution. A solution that does not dissolve is  
 (5) \_\_\_\_\_ insoluble. (6) \_\_\_\_\_ liquid \_\_\_\_\_ solution is the most common  
 type of solution. If one liquid is soluble in another liquid, such as water and oil, the  
 two liquids are (7) \_\_\_\_\_ miscible. However, if one liquid is insoluble in another  
 the liquids are (8) \_\_\_\_\_ immiscible.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.



1. Hydration is solvation in which the solute is a solid. What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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Oxidation, Matter and Change • Chapter 11

Name \_\_\_\_\_

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**Section 15.2** **solvent****Section 15.3** **solvent**

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a test tube with the test of solution, stirring the solution, breaking the solution into smaller pieces, and heating the solution.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with their graphs.

True \_\_\_\_\_ 16. The overall energy change that occurs when a solute dissolves is heat of solution.

False \_\_\_\_\_ 17. Solubility is a property of the solute. Amount of solute that dissolves in a given amount of solvent is a property of temperature and pressure.

False \_\_\_\_\_ 18. Solubility continues to change as the solution heat is lost in the cooling process.

True \_\_\_\_\_ 19. In a saturated solution, solute and crystallization are in equilibrium.

True \_\_\_\_\_ 20. Additional solute can be dissolved in an unsaturated solution.

True \_\_\_\_\_ 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Study Guide For Content

**Ruth Swan**

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Worksheets allow for differentiated instruction through a wide range of question formats There are worksheets and study

tools for each section of the text that help teachers track students progress toward understanding concepts Guided Reading

Activities help students identify and comprehend the important information in each chapter

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