

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

### Solutions

#### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solute and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air's (4) soluble substance. A substance that does not dissolve is (5) insoluble. (6) Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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## CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

#### Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

True \_\_\_\_\_ 14. The overall energy change that occurs when a solution forms is called the heat of solution.

False \_\_\_\_\_ 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False \_\_\_\_\_ 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

True \_\_\_\_\_ 17. In a saturated solution, solvation and crystallization are in equilibrium.

True \_\_\_\_\_ 18. Additional solute can be dissolved in an unsaturated solution.

True \_\_\_\_\_ 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Study Guide for Content Mastery

# Study Guide For Content Master Chemistry Key

**Daniela Niemeyer**



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DAT: Dental Admissions Test: Includes 3 Full Length Practice Tests + Online Access to Video Tutorials Barron's Educational Series, Joseph Di Rienzo, John J. Ference, Nicole D. Cornell, Edwin H. Hines, John Swartwood, 2018-05-15 This brand new manual prepares dental school applicants across the United States and Canada to pass the required admissions test. It features Three full length model tests including a diagnostic test. All answers explained in detail. Access to video tutorials from the authors and more. Test takers will also find thorough reviews of all DAT test topics: a general survey of the natural sciences including biology, chemistry, and organic chemistry, as well as testing for perceptual ability, reading comprehension, and quantitative reasoning. **ONLINE PRACTICE TEST** Students will also get access to one additional full length online DAT test with all questions answered and explained. This online exam can be easily accessed by smartphone, tablet, or computer.

*The Reflective Educator's Guide to Practitioner Inquiry* Nancy Fichtman Dana, Diane Yendol-Hoppey, Logan Rutten, 2025-05-30 Transform your teaching and shape education through the power of inquiry. Grounded in real world examples and more than 30 years of research in professional development, the fifth edition of *The Reflective Educator's Guide to Practitioner Inquiry* addresses how inquiry fosters curiosity, reflection, and practical action to enhance effective classroom learning. This latest edition invites educators to view inquiry as a process, a product, and a stance. The book offers new sections on the relationship between AI and teacher inquiry and the ways in which inquiry is changing with the times. Additional features and updates include: A new chapter that provides a step by step guide to crafting an intentional, actionable research plan. Access exercises, inductive and deductive data analysis worksheets, ethical guidelines, and examples addressing today's issues in education. A discussion of the connections between practitioner inquiry and supporting the academic success of every student. *The Reflective Educator's Guide to Practitioner Inquiry* fifth edition empowers educators at every stage in their careers to investigate their practice, drive sustained professional growth, and harness inquiry's potential to create classrooms where both students and teachers thrive.

**Organic Chemistry** L. G. Wade, 2006 Survival Handbook for the New Chemistry Instructor Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching. Serving as a reference to answer specific questions, new chemistry faculty encounter, this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls. It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course. Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching. **Study Guide**

**Essential Biology with Physiology** Edward J. Zalisko, 2003-07 Students can master key concepts and earn a better grade with the thought provoking exercises found in this study guide. Study advice, tables, quizzes, and crossword puzzles help students test their understanding of biology. The Study Guide also includes references to student media activities on the

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Smoot,Smoot,Richard G. Smith,Jack Price,1998 **The Journal of Education** ,1904 **Curriculum Review** ,1985  
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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Study Guide For Content Master Chemistry Key**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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