

Covalent Bonding

Section 11 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the sentence.

covalent bond	valence	sigma bond	double bond	p bond
When sharing of electrons occurs, the electrons between atoms that results is called	(a) covalent bond	When each of the atoms is involved in electron sharing		

(a) **covalent bond** When each of the atoms is involved in electron sharing energy is released, and the process is (b) **exothermic**. When two or more atoms bond by means of electrons sharing the resulting particle is called an (b) **molecule**. If the electrons shared are located between the two atoms, the covalent bond is called an (b) **sigma bond**. If the electrons shared are involving electrons of parallel orbitals, the covalent bond is called an (b) **p bond**.

In your textbook, read about ionic and covalent bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

1. In which two elements is it a hydrogen, nitrogen, and oxygen atom(s) are? (a) a single atom (b) a molecule containing two atoms (c) two atoms (d) a molecule containing four atoms

2. The energy electrons are shared in a double covalent bond? (a) one (b) two (c) three (d) four

3. Bond length is the distance between (a) two molecules of the same substance (b) the electrons in two different atoms (c) the following relationship relating to bond length is generally correct? (d) the shorter the bond, the stronger the bond (e) the shorter the bond, the weaker the bond (f) the shorter the bond, the stronger the bond (g) the shorter the bond, the longer the bond (h) the longer the bond

Section 12 Naming Molecules

In your textbook, read about how binary compounds and salts are named from their formulas.

(a) For each statement below, write true or false.

(b) _____ 1. Binary molecular compounds are generally composed of a metal and a nonmetal.

(c) _____ 2. The second element in the formula of a binary compound is always a metal.

(d) _____ 3. The first two letters in the name of a binary compound are always a metal.

(e) _____ 4. The prefix two indicates two atoms.

(f) _____ 5. The prefix two indicates two atoms.

(g) _____ 6. In naming the first element in a formula, the prefix names is not used.

(h) _____ 7. For binary acids, the hydrogen part of the compound is named using the prefix hydro-.

(i) _____ 8. A compound contains only two elements.

(j) _____ 9. If the name of the second element in a binary salt is not the acid name, then the suffix -ate is used.

In your textbook, read about naming organic compounds and isomers.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
1. H_2O	a. hydrochloric acid
2. H_2SO_4	b. aluminum bromide
3. Na_2O	c. carbon monoxide
4. Na_2SO_4	d. nitric acid
5. NaOH	e. ammonia
6. H_2CO_3	f. zinc acid
7. H_2SO_3	g. carbon acid
8. H_2O_2	h. sulfuric acid

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Daniel C. Pantaleo

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