

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

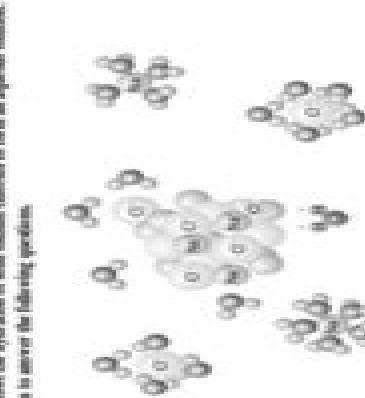
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solute _____ is dissolved in a solvent, or (5) _____ is in (6) _____ solution. A substance that does not dissolve is (7) _____ insoluble _____ in (8) _____ liquid _____ solution or the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, the two liquids are (9) _____ miscible _____ because if one liquid is insoluble in another, the liquids are (10) _____ immiscible _____.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.



1. Indicate in solute in which the solute is in a (a) (b) (c) (d)

form a solution.

Study Guide for Content Mastery

Section 15.2 Solubility

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CHAPTER 15 STUDY GUIDE FOR CONSTITUTANT MASTERY**Section 15.1** *continued*

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the interaction of the water molecules around the sodium ion and chloride ion.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a test tube with the test of solution may be dissolved.

Starting at shaking the solution, breaking the solution into smaller pieces, and heating the solution.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with type of solute.

True _____ 16. The overall energy change that occurs when a solute dissolves in a solvent is called the heat of solution.

False _____ 17. Solubility is a property of the solute, amount of either the solute or a given amount of solute in a specified temperature and pressure.

False _____ 18. Solubility continues to drop as the solution has a high heat of solution.

True _____ 19. In a saturated solution, solute and solvents are in equilibrium.

True _____ 20. Additional solute can be dissolved in an unsaturated solution if the solution becomes a supersaturated solution.

True _____ 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

Section 15.2 *continued***Study Guide for Content Mastery**

Study Guide For Content Mastery

Alton Biggs

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Chemistry: Matter & Change, Study Guide For Content Mastery, Student Edition McGraw Hill,2001-03-30 Study Guide and Reinforcement Worksheets allow for differentiated instruction through a wide range of question formats There are worksheets and study tools for each section of the text that help teachers track students progress toward understanding concepts Guided Reading Activities help students identify and comprehend the important information in each chapter

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