

Physics Study Guide 3

- Heat

- **Convert different temperature scales from one to another.**

Fahrenheit to Celsius: $(32^{\circ}\text{F} - 32) \times 5/9 = 0^{\circ}\text{C}$

Example: 56 Fahrenheit to ? C

$$(\underline{56^{\circ}\text{F}} - 32) \times 5/9 = \underline{13.33^{\circ}\text{C}}$$

Celsius to Kelvin: $0^{\circ}\text{C} + 273 = 273\text{K}$

Example: 56 Celsius to ? K

$$\underline{56^{\circ}\text{C}} + 273 = \underline{329\text{K}}$$

Entropy - Whenever energy freely transforms from one form to another, the direction of transformation is toward a state of greater disorder and therefore toward one of greater entropy

- **Describe thermal expansion and how it relates to temperature.**

Thermal expansion - where an object expands and becomes larger due to a change in its temperature.

higher temperature means that the molecules are moving faster on average, causing them to take up more space.

thermal energy in a system can move from lower to higher temperatures, when external work is done on the system.

- **Define the ideal gas law and describe how it relates pressure, volume, and temperature**

Enthalpy - The change in thermal energy of a system plus the product of its pressure and volume

The term ideal gas refers to a hypothetical gas composed of molecules that: do not attract or repel each other, and themselves (the molecules) take up no volume.

Ideal Gas Law -

- **Define heat and relate the heat transferred in materials to temperature.**

Heat - thermal energy that flows from a substance of higher temperature to a substance of lower temperature.

Heat \neq thermal energy \rightarrow Heat is energy in transit

Energy transfers are from objects at higher temperatures to lower temperatures

Study Guide Heat For Physics

Donald R. Pitts, Leighton E. Sissom



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