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web feb 13 2023 we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web consider the thermochemical equation below $4 \text{NO} + \text{O}_2 \rightarrow 4 \text{NO}_2$ $\Delta H = 566 \text{ kJ}$ is this reaction endothermic or exothermic endothermic $\Delta H < 0$ b how much heat will be absorbed or released if 10 g of NO is reacted with 10 g O₂

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web sep 5 2023 3 carry out a thermochemical calculation on all these species at their optimized geometries 4 sum up the enthalpies free energy of the products and the reagents keeping into account the

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web answer a 1 explanation the entire reaction was multiplied by 2 therefore ΔH must also be halved 3 answer c explanation option c is incorrect because 566 kJ/mol is the ΔH per mole of O₂ and not for 2 moles even though options b and d are expressed in grams they are still correct because 56

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web q 140 g 4 184 j 16oc g oc q 9372 16 j 35 5 g x 1 mole kcn 0 5453 moles 65 1 g 0 5453 mole moles 5 when a 15 g sample of srcl2 dissolves in 80 g of water in a calorimeter the temperature drops from 42 oc to 15 c calculate the Δh for the process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies read more

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