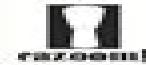


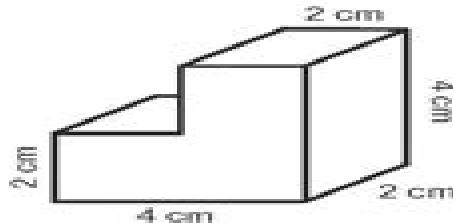
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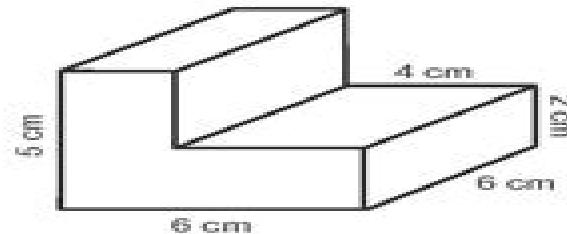
Find the volume of the following objects.

Section A

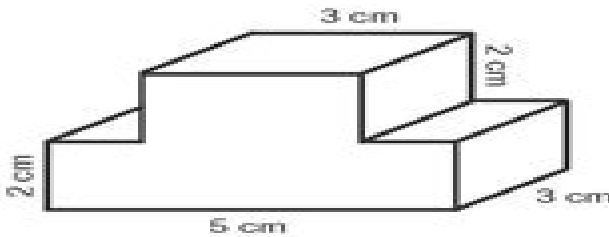
1)



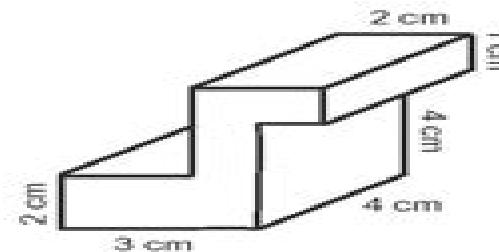
2)



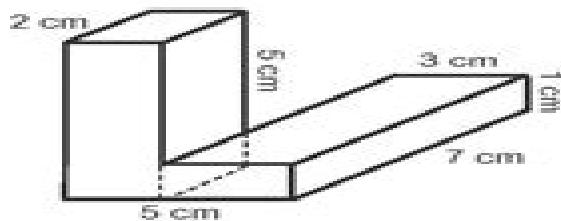
3)



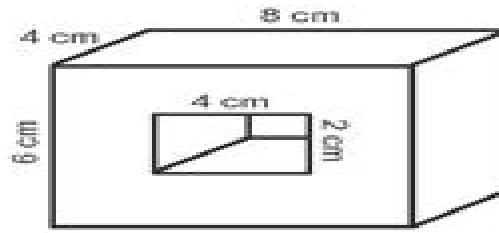
4)



5)



6)



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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant k is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of k

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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we use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $Ca(OH)_2 \rightarrow Ca^{2+} + 2OH^-$ 5 $\times 10^{-2}$ M calcium ions 5 $\times 10^{-2}$ M 2.0 M hydroxide ions calculate calcium hydroxide K_{sp} using the concentrations of calcium ions and hydroxide ions

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web question experiment 8 determining K_{sp} for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 HCl molarity 0.050 M volume of filtered $Ca(OH)_2$ solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml OH in saturated $Ca(OH)_2$ M $Ca(OH)_2$ M K_{sp} of $Ca(OH)_2$ average K_{sp} of $Ca(OH)_2$

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web the chemical equation used was $Ca(OH)_2 + 2HCl \rightarrow CaCl_2 + 2H_2O$ the solubility product was determined to be $K_{sp} = Ca^{2+} \cdot OH^-$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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