

Topic 7.7: Calculating Equilibrium Concentrations

- The concentrations or partial pressures of species at equilibrium can be predicted given the balanced equation, initial concentrations, and the appropriate K .

EXAMPLE

Consider the reaction represented by the equation:



- Trial #1:
6.00 M $\text{Fe}^{3+}(\text{aq})$ and 10.0 M $\text{SCN}^{-}(\text{aq})$ are mixed at a certain temperature and at equilibrium the concentration of $\text{FeSCN}^{2+}(\text{aq})$ is 4.00 M.

What is the value for the equilibrium constant for this reaction?

	Fe^{3+}	+	SCN^{-}	$\rightleftharpoons$	FeSCN^{2+}
Initial					
Change					
Equilibrium					

- Set up the ICE table with the given values.
- Fill in the changes.
- Calculate the equilibrium values.
- Plug Eq values into the K expression.

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