

**STUDY GUIDE AP Chemistry**  
**CHAPTER NINE- MOLECULAR GEOMETRY AND BONDING THEORIES**  
**SECTIONS 9.1 THROUGH 9.6 ONLY**

### 9.1 Molecular Shapes

- Lewis structures give atomic connectivity: they tell us which atoms are physically connected to which atoms.
- The shape of a molecule is determined by its **bond angles**.
  - The angles made by the lines joining the nuclei of the atoms in a molecule are the bond angles.
- Consider  $\text{CCl}_4$ :
  - Experimentally we find all  $\text{Cl}-\text{C}-\text{Cl}$  bond angles are  $109.5^\circ$ .
  - Therefore, the molecule cannot be planar.
  - All Cl atoms are located at the vertices of a tetrahedron with the C at its center.
- In order to predict molecular shape, we assume that the valence electrons repel each other.
  - Therefore, the molecule adopts the three-dimensional geometry that minimizes this repulsion.
  - We call this model the Valence Shell Electron Pair Repulsion (VSEPR) model.

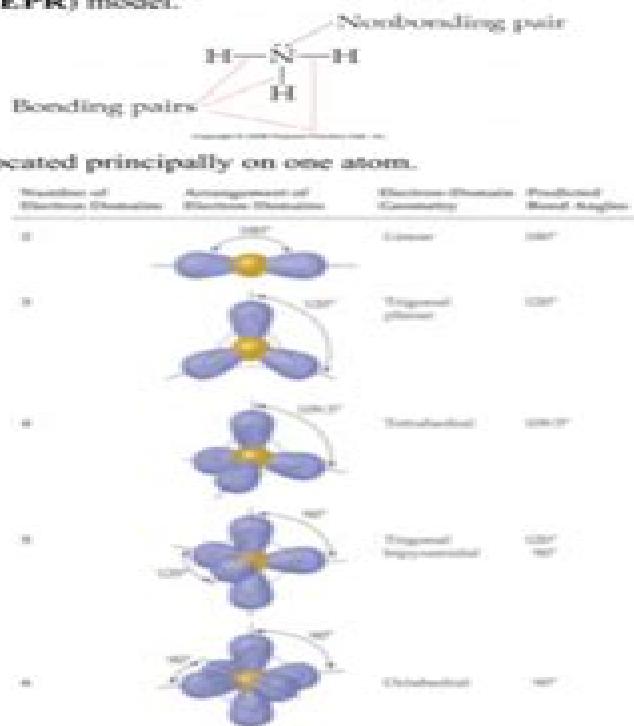
### 9.2 The VSEPR Model

- A covalent bond forms between two atoms when a pair of electrons occupies the space between the atoms.
  - This is a **bonding pair** of electrons.
  - Such a region is an **electron domain**.
- A **nonbonding pair** or **lone pair** of electrons defines an electron domain located principally on one atom.
- Example:  $\text{NH}_3$  has three bonding pairs and one lone pair.
- VSEPR predicts that the best arrangement of electron domains is the one that minimizes the repulsions among them.
  - The arrangement of electron domains about the central atom of a molecule is its **electron-domain geometry**.
  - There are five different electron-domain geometries:
    - linear (two electron domains), trigonal planar (three domains), tetrahedral (four domains), trigonal bipyramidal (five domains) and octahedral (six domains).
- The **molecular geometry** is the arrangement of the atoms in space.
- To determine the shape of a molecule we must distinguish between lone pairs and bonding pairs.
  - We use the electron-domain geometry to help us predict the molecular geometry.
    - Draw the Lewis structure.
    - Count the total number of electron domains around the central atom.
    - Arrange the electron domains in one of the five geometries to minimize electron-electron repulsion.
    - Next, determine the 3-D structure of the molecule.
      - We ignore lone pairs in the molecular geometry.
      - Describe the molecular geometry in terms of the bonded atoms.
    - Multiple bonds are counted as one electron domain.

#### The Effect of Nonbonding Electrons and Multiple Bonds on Bond Angles

We refine VSEPR to predict and explain slight distortions from "ideal" geometries.

- Consider three molecules with tetrahedral electron domain geometries:  $\text{CH}_4$ ,  $\text{NH}_3$ , and  $\text{H}_2\text{O}$ .
  - By experiment, the  $\text{H}-\text{X}-\text{H}$  bond angle decreases from C ( $109.5^\circ$  in  $\text{CH}_4$ ) to N ( $107^\circ$  in  $\text{NH}_3$ ) to O ( $104.5^\circ$  in  $\text{H}_2\text{O}$ ).
    - A bonding pair of electrons is attracted by two nuclei. They do not repel as much as lone pairs which are primarily attracted by only one nucleus.
  - Electron domains for nonbonding electron pairs thus exert greater repulsive forces on adjacent electron domains.
    - They tend to compress the bond angles.
    - The bond angle decreases as the number of nonbonding pairs increases.
  - Similarly, electrons in multiple bonds repel more than electrons in single bonds (e.g., in  $\text{Cl}_2\text{CO}$  the  $\text{O}-\text{C}-\text{Cl}$  angle is  $124.3^\circ$ , and the  $\text{Cl}-\text{C}-\text{Cl}$  bond angle is  $111.4^\circ$ ).



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**Margaret G. Kimble**

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