

Chemistry Study Guide

Periodic Table of the Elements

¹ ₁ H																	² ₂ He																														
³ ₃ Li	⁴ ₄ Be																	⁵ ₅ B	⁶ ₆ C	⁷ ₇ N	⁸ ₈ O	⁹ ₉ F	¹⁰ ₁₀ Ne																								
¹¹ ₁₁ Na	¹² ₁₂ Mg																	¹³ ₁₃ Al	¹⁴ ₁₄ Si	¹⁵ ₁₅ P	¹⁶ ₁₆ S	¹⁷ ₁₇ Cl	¹⁸ ₁₈ Ar																								
¹⁹ ₁₉ K	²⁰ ₂₀ Ca	²¹ ₂₁ Sc	²² ₂₂ Ti	²³ ₂₃ V	²⁴ ₂₄ Cr	²⁵ ₂₅ Mn	²⁶ ₂₆ Fe	²⁷ ₂₇ Co	²⁸ ₂₈ Ni	²⁹ ₂₉ Cu	³⁰ ₃₀ Zn	³¹ ₃₁ Ga	³² ₃₂ Ge	³³ ₃₃ As	³⁴ ₃₄ Se	³⁵ ₃₅ Br	³⁶ ₃₆ Kr																														
³⁷ ₃₇ Rb	³⁸ ₃₈ Sr	³⁹ ₃₉ Y	⁴⁰ ₄₀ Zr	⁴¹ ₄₁ Nb	⁴² ₄₂ Mo	⁴³ ₄₃ Tc	⁴⁴ ₄₄ Ru	⁴⁵ ₄₅ Rh	⁴⁶ ₄₆ Pd	⁴⁷ ₄₇ Ag	⁴⁸ ₄₈ Cd	⁴⁹ ₄₉ In	⁵⁰ ₅₀ Sn	⁵¹ ₅₁ Sb	⁵² ₅₂ Te	⁵³ ₅₃ I	⁵⁴ ₅₄ Xe																														
⁵⁵ ₅₅ Cs	⁵⁶ ₅₆ Ba		⁵⁷ ₅₇ Hf	⁵⁸ ₅₈ Ta	⁵⁹ ₅₉ W	⁶⁰ ₆₀ Re	⁶¹ ₆₁ Os	⁶² ₆₂ Ir	⁶³ ₆₃ Pt	⁶⁴ ₆₄ Au	⁶⁵ ₆₅ Hg	⁶⁶ ₆₆ Tl	⁶⁷ ₆₇ Pb	⁶⁸ ₆₈ Bi	⁶⁹ ₆₉ Po	⁷⁰ ₇₀ At	⁷¹ ₇₁ Rn																														
⁷² ₇₂ Fr	⁷³ ₇₃ Ra		⁷⁴ ₇₄ Rf	⁷⁵ ₇₅ Db	⁷⁶ ₇₆ Sg	⁷⁷ ₇₇ Bh	⁷⁸ ₇₈ Hs	⁷⁹ ₇₉ Mt	⁸⁰ ₈₀ Ds	⁸¹ ₈₁ Rg	⁸² ₈₂ Cn	⁸³ ₈₃ Uut	⁸⁴ ₈₄ Fl	⁸⁵ ₈₅ Uup	⁸⁶ ₈₆ Lv	⁸⁷ ₈₇ Uus	⁸⁸ ₈₈ Uuo																														
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

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Paul A. Bartlett, Judith G. Koch



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