

Chapter 19 Oxidation-Reduction Reactions

Chapter Opener

— **Chapter Overview, TE** Review the objectives listed in the Student Edition.

Section 1 Oxidation and Reduction

PACING: 45 minutes

PENNSYLVANIA ACADEMIC STANDARDS FOR SCIENCE AND TECHNOLOGY:
3.1.E.1 Describe how fundamental science and technology concepts are used to solve practical problems.

Objectives

1. Assign oxidation numbers to reactant and product species.
2. Define oxidation and reduction.
3. Explain what an oxidation-reduction (redox reaction) is.

Focus (5 minutes)

— **Lesson Starter, TE** Review oxidation numbers and have students familiarize themselves with the rules for assigning oxidation numbers in **Table 1**.

MOTIVATE (10 minutes)

— **Demonstration, A Spontaneous Redox Reaction, TE** Demonstrate a redox reaction by using KI, KMnO_4 , and HCl and make sure students can explain the formation of the precipitate and the color changes.

TEACH (20 minutes)

PowerPoint

- **Visual Strategy, Figure 1, TE** Make sure students understand how this figure illustrates how an element can exist in a number of different oxidation states. (GENERAL)
- **Chemistry in Action, Photochromic Lenses, SE** Photochromic eyeglasses darken and become transparent according to ultraviolet light because of an oxidation-reduction reaction.
- **Application, TE** Sodium hypochlorite, NaOCl , is the chemical name for household bleach used to remove stains from clothing through a redox reaction.
- **Discussion, TE** Students should follow up the spontaneous redox reaction demonstration by explaining the results in terms of oxidation and reduction and by writing formulas, using **Table 1** for the oxidation numbers.
- **Chemistry in Action, Skunk-Spray Remedy, SE** Paul Krebaum invented a better way than tomato juice to get rid of skunk spray odor.

Study Guide Redox Reactions Answers

Julie R. Frentrup

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