

Chapter 19 Oxidation-Reduction Reactions

Chapter Opener

__ **Chapter Overview, TE** Review the objectives listed in the Student Edition.

Section 1 Oxidation and Reduction

PACING: 45 minutes

PENNSYLVANIA ACADEMIC STANDARDS FOR SCIENCE AND TECHNOLOGY:

3.1.E.1 Describe how fundamental science and technology concepts are used to solve practical problems.

Objectives

1. Assign oxidation numbers to reactant and product species.
2. Define *oxidation* and *reduction*.
3. Explain what an oxidation-reduction (redox reaction) is.

Focus (5 minutes)

__ **Lesson Starter, TE** Review oxidation numbers and have students familiarize themselves with the rules for assigning oxidation numbers in **Table I**.

MOTIVATE (10 minutes)

__ **Demonstration, A Spontaneous Redox Reaction, TE** Demonstrate a redox reaction by using KI, KMnO_4 , and HCl and make sure students can explain the formation of the precipitate and the color changes.

TEACH (20 minutes)

__ **PowerPoint**

__ **Visual Strategy, Figure 1, TE** Make sure students understand how this figure illustrates how an element can exist in a number of different oxidation states. (**GENERAL**)

__ **Chemistry in Action, Photochromic Lenses, SE** Photochromic eyeglasses darken and become transparent according to ultraviolet light because of an oxidation-reduction reaction.

__ **Application, TE** Sodium hypochlorite, NaOCl , is the chemical name for household bleach used to remove stains from clothing through a redox reaction.

__ **Discussion, TE** Students should follow up the spontaneous redox reaction demonstration by explaining the results in terms of oxidation and reduction and by writing formulas, using **Table I** for the oxidation numbers.

__ **Chemistry in Action, Skunk-Spray Remedy, SE** Paul Krebaum invented a better way than tomato juice to get rid of skunk spray odor.

Study Guide Redox Reactions Answers

Julie R. Frentrup



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Table of Contents Study Guide Redox Reactions Answers

1. Understanding the eBook Study Guide Redox Reactions Answers
 - The Rise of Digital Reading Study Guide Redox Reactions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Redox Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Redox Reactions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Redox Reactions Answers

- Personalized Recommendations
- Study Guide Redox Reactions Answers User Reviews and Ratings
- Study Guide Redox Reactions Answers and Bestseller Lists
- 5. Accessing Study Guide Redox Reactions Answers Free and Paid eBooks
 - Study Guide Redox Reactions Answers Public Domain eBooks
 - Study Guide Redox Reactions Answers eBook Subscription Services
 - Study Guide Redox Reactions Answers Budget-Friendly Options
- 6. Navigating Study Guide Redox Reactions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Redox Reactions Answers Compatibility with Devices
 - Study Guide Redox Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Redox Reactions Answers
 - Highlighting and Note-Taking Study Guide Redox Reactions Answers
 - Interactive Elements Study Guide Redox Reactions Answers
- 8. Staying Engaged with Study Guide Redox Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Redox Reactions Answers
- 9. Balancing eBooks and Physical Books Study Guide Redox Reactions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Redox Reactions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Redox Reactions Answers
 - Setting Reading Goals Study Guide Redox Reactions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Redox Reactions Answers

- Fact-Checking eBook Content of Study Guide Redox Reactions Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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