

CHAPTER 9 STUDY GUIDE

Chemical Reactions

Section 9.1 Reactions and Equations

In your textbook, read about evidence of chemical reactions.

For each statement, write *yes* if evidence of a chemical reaction is present. Write *no* if there is no evidence of a chemical reaction.

- _____ 1. A tomato smells rotten.
- _____ 2. A drinking glass breaks into smaller pieces.
- _____ 3. A piece of ice melts.
- _____ 4. Drain cleaner is mixed with water and the solution becomes warm.
- _____ 5. Candle wax burns.
- _____ 6. Molten candle wax solidifies.
- _____ 7. Green leaves turn yellow and red as the seasons change.
- _____ 8. Baking powder produces a gas that makes a cake rise.

In your textbook, read about how to represent chemical reactions and how to balance chemical equations.

Use the terms below to complete the passage. Each term may be used once, more than once, or not at all.

arrow	plus sign	(s)	(l)
reactant	product	(g)	(aq)

The fuel for the space shuttle is hydrogen, which burns in oxygen to produce water vapor and energy. In this chemical reaction, hydrogen is a(n) **(9)** _____, oxygen is a(n) **(10)** _____, and water vapor is a(n) **(11)** _____. In a chemical equation for this reaction, a(n) **(12)** _____ is used to separate hydrogen and oxygen from water vapor and energy. A(n) **(13)** _____ is used to separate the symbols for hydrogen and oxygen. A(n) **(14)** _____ symbol is used to tell the state of hydrogen in the reaction, a(n) **(15)** _____ symbol is used for the state of oxygen, and a(n) **(16)** _____ symbol is used for the state of water vapor.

Study Guide Section 2 Chemical Reactions

Thomas Elliott Taylor



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Robinson Chemistry Study Guide Robinson,1992 **Resources in Education** ,1984 **Study Guide for Chemical Principles** Thomas Elliott Taylor,1979 **Basic Concepts of Chemistry, Study Guide** Leo J. Malone,2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems Chemistry of Matter ,1993 As the reader read s this textbook he will learn about the interactions of matter that can occur in a test tube in nature and even inside himself P 9

Chemistry Oxtoby,1994 **A-level Chemistry Challenging Drill Questions (Yellowreef)** Thomas Bond,Chris Hughes,2016-04-07 according to syllabus for exam up to year 2017 completely covers all question types since 2003 full set of step by step solution approaches sold separately answer keys provided provides teachers comments revealing common mistakes wrong habits buy print edition online at www.yellowreef.com to enjoy attractive discounts complete eBook edition and concise eBook edition available also suitable for Cambridge GCE AL H1 H2 Cambridge International AL Cambridge Pre University Books available for other subjects including Physics Chemistry Biology Mathematics Economics English Primary level Secondary level GCE O level GCE A level iGCSE Cambridge A level Hong Kong DSE Concise eBooks are tailored for quick revision whereas Complete eBooks are for detailed studies visit www.yellowreef.com for sample chapters and more

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