

## NUCLEAR CHEMISTRY STUDY GUIDE—CH. 21

1. Write the symbol for uranium—238 in isotope format (with top and bottom #s).
2. What does the top # represent?
3. What does the bottom # represent?
4. What particles are found in the nucleus?
5. Why does the strong nuclear force exist?
6. What is the electrostatic force?
7. What particles does the electrostatic force deal with?
8. Define nuclear reactions.
9. Define nuclear fission and fusion.
10. Nuclear reactions usually produce \_\_\_\_\_ amounts of energy, and isotopes become \_\_\_\_\_ stable. The \_\_\_\_\_ number and the \_\_\_\_\_ number may also change.
11. What is transmutation?
12. Write the beta decay of radium—226.
13. Write the alpha decay of radium—226.
14. Why would the isotopes and the end of problems #13 and #14 still be unstable?
15. Define nuclear radiation.
16. List the 3 radioactive particles in order of INCREASING penetrating power.
17. What is radioactive decay?
18. Alpha particles are basically \_\_\_\_\_ nuclei, and beta particles are basically high energy \_\_\_\_\_.
19. What is half life?
20. What is radioactive dating?
21. What is cobalt—60 used for?
22. Does fission or fusion produce nuclei with heavier mass than the original mass?
23. Name 3 differences between chemical and nuclear reactions.
24. What is carbon—13 radioactive?
25. Elements past atomic # \_\_\_\_\_ are radioactive.
26. If  $1/16$  of a sample of a radioactive isotope is left after nuclear decay, how long did this take if the  $\frac{1}{2}$  life is 4 days?
27. If the half life of a radioactive isotope is 4 days, how much of a 50 gram sample is left after 12 days?

# Study Guide Nuclear Chemistry Answers

**Leo J. Malone**



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**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03  
The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

**Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry** Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15  
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**General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999      *Student Self-study Guide* Donald W. Shive, 1975

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