

## METALS

A metal is an element which has the ability to lose electron(s) and become positively charged. There are about 105 different elements. Of these 84 are metals and 21 are non metals.

### GENERAL PROPERTIES

Metals generally have the following properties.

- i. Have high melting and boiling points
- ii. Are good conductors of heat and electricity
- iii. Have a high density – metal atoms are packed very closely
- iv. Are solids at rtp, except for mercury which is a liquid.
- v. Are malleable i.e. can be beaten into sheets
- vi. Are ductile i.e. can be made into wires.
- vii. Are sonorous i.e. produce sound when beaten
- viii. Are lustrous i.e. can be polished or have a shining surface.
- ix. They are strong under tension and compression. That means they can withstand stretching and crushing without breaking.
- x. React with oxygen to form basic oxides
- xi. Form positively charged ions (Cations) by losing electrons.

### WHY METALS ARE USED AS ALLOYS

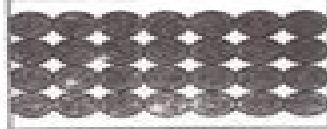
Pure metals are usually too soft and weak for most uses. A gold chain is never made of pure gold, as the chain would be too soft and would break easily. Instead, it is alloyed with other elements to make it more shiny and stronger.

Metals are mainly used as alloys for the following reasons

- To improve on their properties.
- To improve on their appearance.
- To make metals to be less brittle.
- To Make metals strong and harder.

A comparison between pure metals and alloys is shown below.

Pure metal	Alloy
Are usually soft and can easily bend	Are strong and hard, not easily bent
Have higher melting points	Have lower melting points than the melting points of their components
Atoms can slide over each other easily as atoms are of the same size	Atoms cannot slide over each other easily as foreign atoms are of different size and disrupt the orderly arrangement of the metal atoms

  
*Diagrammatic representation of a pure metal*

  
*Diagrammatic representation of an alloy*

# Study Guide Metals Answers

**Will Marchment**

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## **Table of Contents Study Guide Metals Answers**

1. Understanding the eBook Study Guide Metals Answers
  - The Rise of Digital Reading Study Guide Metals Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Metals Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Metals Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Metals Answers
  - Personalized Recommendations
  - Study Guide Metals Answers User Reviews and Ratings
  - Study Guide Metals Answers and Bestseller Lists
5. Accessing Study Guide Metals Answers Free and Paid eBooks
  - Study Guide Metals Answers Public Domain eBooks
  - Study Guide Metals Answers eBook Subscription Services
  - Study Guide Metals Answers Budget-Friendly Options

6. Navigating Study Guide Metals Answers eBook Formats

- ePUB, PDF, MOBI, and More
- Study Guide Metals Answers Compatibility with Devices
- Study Guide Metals Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Study Guide Metals Answers
- Highlighting and Note-Taking Study Guide Metals Answers
- Interactive Elements Study Guide Metals Answers

8. Staying Engaged with Study Guide Metals Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Study Guide Metals Answers

9. Balancing eBooks and Physical Books Study Guide Metals Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Study Guide Metals Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Study Guide Metals Answers

- Setting Reading Goals Study Guide Metals Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Study Guide Metals Answers

- Fact-Checking eBook Content of Study Guide Metals Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements

- Interactive and Gamified eBooks

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