

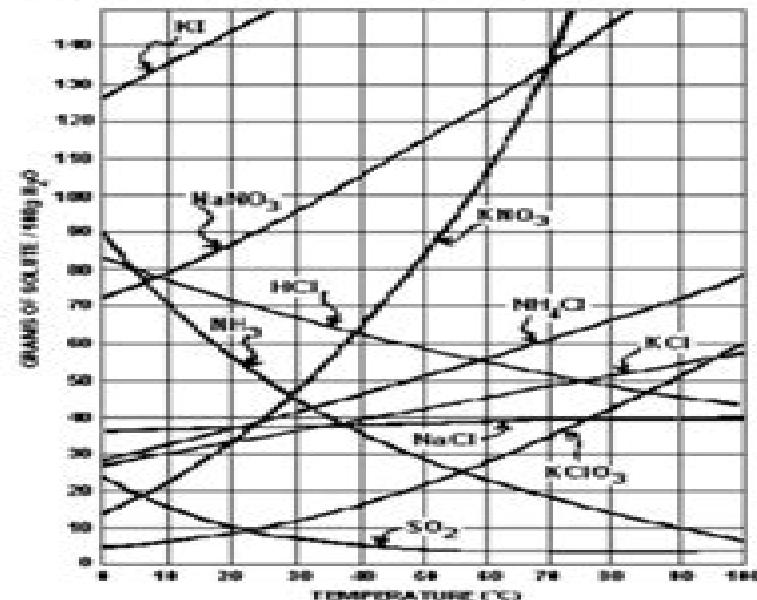
### Solutions Level II

**Section I: Reading Solubility Curves – Use the solubility curve below to answer the following questions.**

- What relationship exists between solubility and temperature for the ionic substances shown?
- What relationship exists between solubility and temperature for the only gas,  $\text{SO}_2$ , on the graph?

Indicate if the following solutions are saturated, unsaturated or supersaturated.

- 20 grams of  $\text{NH}_3$  completely dissolved in 100 grams of water at  $80^\circ\text{C}$ .
- 80 grams of  $\text{KNO}_3$  completely dissolved in 100 grams of water at  $55^\circ\text{C}$ .
- 112 grams of  $\text{NaNO}_3$  completely dissolved in 100 grams water at  $45^\circ\text{C}$ ?
- 10 grams of  $\text{KClO}_3$  completely dissolved in 100 grams water at  $60^\circ\text{C}$ ?
- 30g  $\text{KCl}$  in 100g water at  $10^\circ\text{C}$ ?



**Section II: Electrolytes – Complete the following table. The first one has been done for you as an example.**

	Formula?	Ionic, Acid/Base, or Covalent?	Soluble in water?	Electrolyte?
Sodium chloride	$\text{NaCl}$	Ionic	Yes	Yes
1. Acetic Acid				
2. Silver bromide				
3. Hydrofluoric acid				
4. Carbon dioxide				
5. Strontium hydroxide				
6. Magnesium chlorate				
7. Ammonium nitrate				

# Study Guide Solubility Graph Answers

**Leo J. Malone**

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