

The Miller Group's Reported Production: KEY

1. Identify the actual value of the number of representative particles in each of the following quantities:

a. 2.50 moles of gold

$$\frac{2.50 \text{ mol Au}}{1 \text{ mol}} \times \frac{6.02 \times 10^{23} \text{ particles/mol}}{1 \text{ mole}} = 1.51 \times 10^{24} \text{ atoms Au}$$

b. 10.50 moles of nitrogen gas

$$\frac{10.50 \text{ mol N}_2}{1 \text{ mol}} \times \frac{6.02 \times 10^{23} \text{ molecules N}_2}{1 \text{ mole N}_2} = 6.32 \times 10^{24} \text{ molecules N}_2$$

c. 0.750 moles of potassium permanganate

$$\frac{0.750 \text{ mol KMnO}_4}{1 \text{ mol}} \times \frac{6.02 \times 10^{23} \text{ formula units KMnO}_4}{1 \text{ mole KMnO}_4} = 4.52 \times 10^{23} \text{ formula units KMnO}_4$$

2. Calculate the number of moles of the substance that contains the following number of representative particles:

a. 8.800 or 8.80×10^{24} atoms of sodium

$$\frac{8.80 \times 10^{24} \text{ atoms Na}}{1 \text{ mol}} \times \frac{1 \text{ mole Na}}{6.02 \times 10^{23} \text{ atoms Na}} = 14.6 \text{ mol Na}$$

b. 6.50 or 6.50×10^{24} molecules of carbon monoxide

$$\frac{6.50 \times 10^{24} \text{ molecules CO}}{1 \text{ mol}} \times \frac{1 \text{ mole CO}}{6.02 \times 10^{23} \text{ molecules CO}} = 10.8 \text{ mol CO}$$

c. 2.400 or 2.40×10^{24} formula units of potassium chloride

$$\frac{2.40 \times 10^{24} \text{ formula units KCl}}{1 \text{ mol}} \times \frac{1 \text{ mole KCl}}{6.02 \times 10^{23} \text{ formula units KCl}} = 3.99 \text{ mol KCl}$$

Study Guide The Mole Supplemental Problems

Jin-Ying Zhang



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