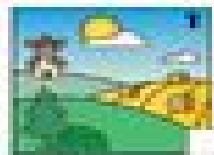


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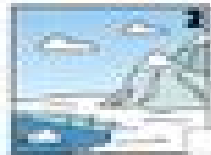
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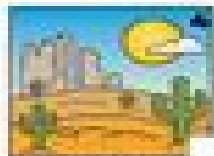
$\text{CO}_2(g) + \text{H}_2\text{O}(l) \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(aq) + \text{O}_2(g)$

What mass of CO_2 is needed to produce 100g of $\text{C}_6\text{H}_{12}\text{O}_6$?



$\text{Fe}_2\text{O}_3(s) + \text{Al}(s) \rightarrow \text{Fe}(l) + \text{Al}_2\text{O}_3(s)$

What mass of Fe can be produced from 100g of Fe_2O_3 ?



$\text{H}_2\text{SO}_4(aq) + \text{NaOH}(aq) \rightarrow \text{Na}_2\text{SO}_4(aq) + \text{H}_2\text{O}(l)$

What mass of NaOH can be produced from 100g of H_2SO_4 ?



$\text{C}_6\text{H}_{12}\text{O}_6(aq) + \text{O}_2(g) \rightarrow \text{C}_6\text{H}_{10}\text{O}_5(s) + \text{H}_2\text{O}(l)$

What mass of H_2O can be produced from 100g of $\text{C}_6\text{H}_{12}\text{O}_6$?

Stoichiometry - Reference Sheet -

Conversion Factors

Steps for Calculations

- 1. BALANCE:** Balance the chemical equation for the reaction.
- 2. LABEL "G" & "L":** Label the equation with the substance that is given with "G" and the substance that is to be unknown with "L".
 Example: $\text{C}_6\text{H}_{12}\text{O}_6(aq) + \text{O}_2(g) \rightarrow \text{C}_6\text{H}_{10}\text{O}_5(s) + \text{H}_2\text{O}(l)$
 Label: $\text{C}_6\text{H}_{12}\text{O}_6(aq)$ is given and $\text{O}_2(g)$ is to be unknown.
- 3. CHOOSE CONVERSION FACTORS:** Choose the conversion factors that will convert the given substance to the unknown substance. (Choose from your conversion table to solve, making sure the units cancel out and the units of the unknown substance are left.)
- 4. SET UP THE CALCULATION:** Write your conversion factors in order on the following template. Include division.

$$\text{Mass of substance (g)} \times \left(\frac{\text{molar mass (g/mol)}}{\text{molar mass (g/mol)}} \right) \times \left(\frac{\text{molar ratio (mol)}}{\text{molar ratio (mol)}} \right) = \text{mass of substance (g)}$$
- 5. CALCULATE:** Solve the equation to find the mass of the unknown substance. Round the answer to the correct number of significant figures.

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- Easy method to teach stoichiometry
- Self-correcting games



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10th European Conference on Games Based Learning , **Chemistry Games** Gregory Gebhardt,2012-10-01 This book is a science education text It is a collection of Chemistry games that teaches the periodic table stoichiometry chemical reactions and the law of conservation of mass **Resources in Education** ,1993 *CHEMISTRY* SILBERBERG,2003

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Proceedings of the ... International Conference on Coordination Chemistry ,1964 **The software catalog microcomputers** Menu (Firm) (Fort Collins, Colo.),1989 Resources in Education ,1993 *Carolina Science and Math* Carolina Biological Supply Company,2003

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- Chapter 2: Essential Elements of Studying Games For Chemistry Stoichiometry
- Chapter 3: Studying Games For Chemistry Stoichiometry in Everyday Life
- Chapter 4: Studying Games For Chemistry Stoichiometry in Specific Contexts
- Chapter 5: Conclusion

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4. In chapter 3, this book will examine the practical applications of Studying Games For Chemistry Stoichiometry in daily life. The third chapter will showcase real-world examples of how Studying Games For Chemistry Stoichiometry can be effectively utilized in everyday scenarios.
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6. In chapter 5, this book will draw a conclusion about Studying Games For Chemistry Stoichiometry. This chapter will summarize the key points that have been discussed throughout the book.

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