

STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

STOICHIOMETRY

Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.

- Stoichiometry is possible because of the Law of Conservation of Mass.
- The law of conservation of mass states that the amount of matter before a reaction occurs is equal to the amount of matter present after the reaction occurs.
- Thanks to this idea, we can build ratios between reactants and product that allow us to calculate amounts of each.

Using Ratios

Reactants and products are related through stoichiometry.

Each coefficient in the balanced equation can be used to define a ratio between the individual reactants and products of the chemical reaction. For the next example, let's use only moles of chloride.

Consider the reaction:

$$2\text{NaCl} + \text{MgSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{MgCl}_2$$

We can write down ratios:

moles of NaCl	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}}{1\text{MgSO}_4}$
moles of NaCl	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}}{2\text{MgCl}_2}$
moles of NaCl	$\frac{2\text{NaCl}}{2\text{NaCl}}$	$\frac{2\text{NaCl}}{1\text{Na}_2\text{SO}_4}$
moles of MgSO ₄	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$	$\frac{1\text{MgSO}_4}{1\text{MgCl}_2}$
moles of MgSO ₄	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$	$\frac{1\text{MgSO}_4}{1\text{Na}_2\text{SO}_4}$
moles of MgSO ₄	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$	$\frac{1\text{MgSO}_4}{2\text{MgCl}_2}$

Example: Finding the Limiting Reactant in the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

5.37 grams of NaCl react with 5.63 grams of Mg. What is the limiting reactant and what mass of MgSO₄ can be formed?

First we need to write $\frac{\text{moles}}{\text{moles}}$ ratios: $\frac{2\text{NaCl}}{2\text{NaCl}}$ and $\frac{1\text{Mg}}{1\text{Mg}}$

Now we multiply them $\frac{2\text{NaCl}}{2\text{NaCl}} \times \frac{1\text{Mg}}{1\text{Mg}}$ to get $\frac{2\text{NaCl}}{1\text{Mg}}$

$$5.37\text{g NaCl} \times \frac{1\text{Mg}}{2\text{NaCl}} = 0.046\text{Mg}$$

$$5.63\text{g Mg} \times \frac{1\text{Mg}}{1\text{Mg}} = 0.046\text{Mg}$$

0.046 grams of MgSO₄ will form and NaCl is the limiting reactant.

Using Mole Ratios

From the last page, we looked at the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

We also defined ratios of moles of each substance using the coefficients of the balanced equation.

Now let's use the ratio we defined:

How many moles of Mg can be produced from 4 moles of NaCl in the reaction above?

We know the ratio of moles of Mg to moles of NaCl according to the balanced equation is 1:2. By writing this ratio into a fraction, we get:

$$\frac{1\text{Mg}}{2\text{NaCl}}$$

and we get:

$$4\text{NaCl} \times \frac{1\text{Mg}}{2\text{NaCl}} = 2\text{Mg}$$

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