

ATOMIC STRUCTURE

RUTHERFORD'S MODEL

- Discovered that an atom is mostly empty space with tiny particles.
- Charged particles move in a small elliptical path around a central nucleus.
- This can only hold the negatively charged particles in the center.

BOHR'S MODEL

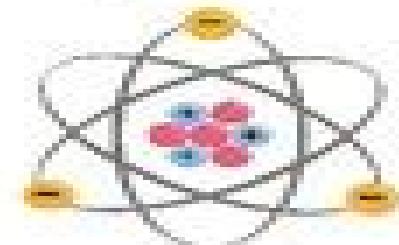
- Shows the elliptical path of an electron.
- Electrons
- Electrons

THE JELLY PUDDING MODEL



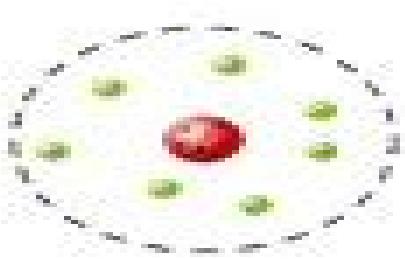
WAVE-MECHANICAL MODEL

- Shows electrons moving in paths, like tiny clouds like in clouds. Orbitals are regions of space likely where an electron.



WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
 - PROTONS - positively charged particles
 - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
 - Electrons DO move around the nucleus in different patterns.



Subatomic Particle	Location in the Atom	Charge	Weight
proton	center (in the nucleus)	positive	1.67 x 10 ⁻²⁴
neutron	center (in the nucleus)	neutral (no charge)	1.67 x 10 ⁻²⁴
electron	around the nucleus in elliptical orbits	negative	1.67 x 10 ⁻²⁴ x 10 ⁻³⁰

IONS

- If an atom gains or loses an electron, it becomes an ion.
- This gives the atom a charge:
 - Loss of electrons, the atom becomes positive.
 - Gains of electrons, the atom becomes negative.

ISOTOPES

- An isotope is when an atom has a different mass number.
- An isotope is still the same element, it is just a little different mass, every atom has all other atoms different in slightly different mass.

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Kenneth W. Raymond

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Table of Contents Study Guide Matter And Atom Answers

1. Understanding the eBook Study Guide Matter And Atom Answers
 - The Rise of Digital Reading Study Guide Matter And Atom Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Matter And Atom Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Study Guide Matter And Atom Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Matter And Atom Answers
 - Personalized Recommendations
 - Study Guide Matter And Atom Answers User Reviews and Ratings
 - Study Guide Matter And Atom Answers and Bestseller Lists

5. Accessing Study Guide Matter And Atom Answers Free and Paid eBooks
 - Study Guide Matter And Atom Answers Public Domain eBooks
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Matter And Atom Answers
 - Highlighting and Note-Taking Study Guide Matter And Atom Answers
 - Interactive Elements Study Guide Matter And Atom Answers
8. Staying Engaged with Study Guide Matter And Atom Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Matter And Atom Answers
9. Balancing eBooks and Physical Books Study Guide Matter And Atom Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Matter And Atom Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Matter And Atom Answers
 - Setting Reading Goals Study Guide Matter And Atom Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Matter And Atom Answers
 - Fact-Checking eBook Content of Study Guide Matter And Atom Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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