

CHEMISTRY Terminology

The ABCs of chemistry

A

A: Symbol for atomic mass number
absolute zero: Lowest temperature; 0 Kelvin
Ac: Symbol for the element actinium, $Z = 89$
accuracy: The closeness of measurement to true value
acetate anion: CH_3COO^- , charge -1
acetic acid: $\text{C}_2\text{H}_4\text{O}_2$
acetone: Dimethyl ketone; organic solvent
acid-base indicator: A chemical that changes color if the pH changes
acid-base titration: A method for determining acid or base concentrations
acid definitions: See Arrhenius acid, Brønsted-Lowry acid, Lewis acid
acid ionization constant (K_{a}): Equilibrium constant for acid dissociation
acidic solution: pH below 7
actinides: Element with $Z = 90$ or above; radioactive
activation energy (E_a): A process's energy barrier
affiliation: Attraction of unlike molecules or materials
Ag: Symbol for the element silver, $Z = 47$
Al: Symbol for the element aluminum, $Z = 13$
alcohol: Organic compound with -OH group; $\text{R}-\text{OH}$
aldehydes: Carbonyl group bonded to hydrogen and an organic group
alkali metal family: Lithium, sodium, potassium, rubidium, cesium, and francium; column #1
alkaline earth metal family: Beryllium, magnesium, calcium, strontium, barium, and radium; column #2
alkane: Hydrocarbon; all C-C single bonds
alkene: Hydrocarbon; 1 or more C=C double bonds
alkyne: Hydrocarbon; 1 or more C≡C triple bonds
allotropes: Two or more forms of an element
alloy: Solution of 2 or more metals
alpha (α): Greek letter that denotes radioactive particle and various scientific variables
alpha particle: Helium nucleus; charge +2
Am: Symbol for the element americium, $Z = 95$
amalgam: Alloy of mercury with other metal
amine: Organic base; R_3N ; modified ammonia
amine acid: Compound with organic acid and organic base property; found proteins and peptides
ammonium: NH_4^+ ; base
ammonium cation: NH_4^+ ; charge +1
ammonium hydroxide: NH_4OH
amphoteric oxide: Exhibits both acid and basic properties
amu: Atomic mass unit; $1/12$ mass of $\text{C}-12$
anion: Ion with a negative charge
anode: Electrode that supports oxidation
antibonding MO: MO is less stable than sigma MOs
AO: Abbreviation for atomic orbital



aq: Aquous; "dissolved in water"
Ar: Symbol for the element argon, $Z = 18$
aromatic: Organic compound with a benzene ring
Arrhenius acid: Produces hydronium ion in water solution
Arrhenius base: Yields hydroxide ion in water solution
Arrhenius equation for rate constant (k):

$$k = A e^{-E_a/RT}$$

E_a = activation energy

arsenide anion: As^{3-} ; charge -3
As: Symbol for the element arsenic, $Z = 33$
At: Symbol for the element astatine, $Z = 85$
atom: Smallest unit in "atmosphere"
atom: Fundamental unit of all matter
atomic mass number (A): Total number of protons and neutrons
atomic number (Z): Number of protons in the nucleus
atomic orbital: Wave motion of electrons in atoms
atomic radius: Empirical measure of atom size
atomic weight: Weighted average of natural isotopes of an element
Au: Symbol for the element gold, $Z = 79$
Aufbau principle: Order the filling of electronic subshells of the elements
Avogadro's law: Volume is proportional to gas moles at fixed pressure and temperature
Avogadro's number: $N_A = 6.022 \times 10^{23}$ mol⁻¹

B: Symbol for the element boron, $Z = 5$
Ba: Symbol for the element barium, $Z = 56$
balanced equations: Each side of the equation has the same number of atoms of each element and charge
balanced redox: Electron loss = electron gain
barium cation: Ba^{2+} ; charge +2
base definition: See Arrhenius base, Brønsted-Lowry base, Lewis base
base ionization constant (K_b): Equilibrium constant for base dissociation
basic solution: pH over 7
battery: Source of electrical power; galvanic cell
Ber: Symbol for the element beryllium, $Z = 4$
benzene: C_6H_6 ; organic solvent; aromatic ring of 6 carbon atoms
beryllium cation: Be^{2+} ; charge +2
beta (β): Greek letter that denotes radioactive particle and a number of scientific variables
beta particles: Energetic electrons from the nucleus; charge -1
Bh: Symbol for the element bismuth, $Z = 107$
Bk: Symbol for the element berkelium, $Z = 97$
bicarbonate anion: HCO_3^- ; charge -1
bimolecular oxygen: O_2
binary compound: Formed from 2 elements
bisulfate anion: HSO_4^- ; charge -1



Bk: Symbol for the element berkelium, $Z = 97$
bleach: Chemical that is a strong oxidizing agent
boiling: Liquid \rightarrow gas at the boiling point
boiling point elevation: A solution has a higher boiling point than a pure solvent
boiling point (T_b): Liquid-gas equilibrium; $P = 1$ atm
boiling point of water: 100°C
bond energy: Energy held by a chemical bond
bond length: Distance between 2 atoms in a bond
bond order for Lewis model: # of bonds divided by # of bonded atoms
bond order for MO treatment: # of filled bonding MOs
bonding electrons: Form chemical bonds in a compound
bonding MO: MO is more stable than AOs
Boyle's law: PV constant for gas; Boyle's law
Br: Symbol for the element bromine, $Z = 35$
bromide anion: Br^- ; charge -1
brønsted-Lowry acid: Proton donor
brønsted-Lowry base: Proton acceptor
buffer: Composed of weak acid and weak base; serves to keep pH constant



C: Symbol for carbon, the SI prefix for "10³"
c: Symbol for the speed of light
°C: Celsius temperature scale
C: Symbol for the element carbon, $Z = 6$
C₂H₄O₂: Acetic acid
Ca: Symbol for the element calcium, $Z = 20$
calcium cation: Ca^{2+} ; charge +2
calcium fluoride: CaF_2
calcium hydroxide: $\text{Ca}(\text{OH})_2$
calories: English unit of heat energy
calorimetry: Study of the release or absorption of heat
carbide anion: C^- ; charge -4
carboxylates: Organic compound; carbon bonded to several -COO groups (sugar and starch)
carbon dioxide: CO_2
carbon monoxide: CO
carbonate anion: CO_3^{2-} ; charge -2
carbonic acid: H_2CO_3
carboxyl group: COOH ; found in carboxylic acids
carboxylic acid: Organic acid; $\text{R}-\text{COOH}$
catalyst: Accelerates a reaction but is not consumed in the reaction; lowers the activation energy
cathode: Electrode that supports reduction
cation: Ion with a positive charge
Cd: Symbol for the element cadmium, $Z = 48$



Study Guide Modern Chemistry

**Peter Atkins, John Krenos, Joseph
Potenza**

Study Guide Modern Chemistry:

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