

### Molecular Polarity Table

A molecule is **non-polar** if all surrounding atoms are same.

A molecule is **polar**, if surrounding atoms are different.

A molecule is **polar**, if central atom has one or more lone pair of electrons.

BP: Bond pair

LP: Lone pair

A: Central atom

X: Surrounding atoms



| Number of electron groups | Type of electron pairs | VSEPR notation | Name of Molecular shape   | Example  | Polarity  |
|---------------------------|------------------------|----------------|---|----------|-----------|
| 2                         | 2 BP                   | $AX_2$         | $X - A - X$<br>linear   | $BeF_2$  | Non-polar |
| 3                         | 3 BP                   | $AX_3$         | <br>trigonal planar        | $BF_3$   | Non-polar |
| 3                         | 2 BP, 1 LP             | $AX_2E$        | <br>angular                | $SnCl_3$ | Polar     |
| 4                         | 4 BP                   | $AX_4$         | <br>tetrahedral            | $CF_4$   | Non-polar |
| 4                         | 3 BP, 1 LP             | $AX_3E$        | <br>trigonal bipyramidal | $PCl_3$  | Polar     |
| 4                         | 2BP, 2LP               | $AX_2E_2$      | <br>angular              | $H_2S$   | Polar     |

# Study Guide Molecular Shape Answer Key

**Arshad Iqbal**

## **Study Guide Molecular Shape Answer Key:**

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