

# The Mole Study Guide

**Mass Composition:** What percent is each element in a compound? What unit do you use?

1) There are 3.75 moles of  $\text{CO}_2$ .

2) There are 1.25 moles of  $\text{H}_2\text{O}$ .

3) There are 0.50 moles of  $\text{CH}_4$ .

4) There are 2.00 moles of  $\text{N}_2$ .

5) There are 0.75 moles of  $\text{O}_2$ .

6) There are 1.50 moles of  $\text{H}_2$ .

7) There are 0.25 moles of  $\text{SO}_2$ .

8) There are 0.50 moles of  $\text{Fe}_2\text{O}_3$ .

9) There are 0.75 moles of  $\text{CaCl}_2$ .

10) There are 1.00 moles of  $\text{NaCl}$ .

11) There are 0.50 moles of  $\text{KNO}_3$ .

12) There are 0.25 moles of  $\text{H}_2\text{SO}_4$ .

13) There are 0.75 moles of  $\text{NH}_3$ .

14) There are 0.50 moles of  $\text{C}_2\text{H}_6$ .

15) There are 0.25 moles of  $\text{H}_2\text{O}_2$ .

16) There are 0.75 moles of  $\text{H}_2\text{O}$ .

17) There are 0.50 moles of  $\text{CO}_2$ .

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## The Mole Study Guide

Fill in the blank boxes for the image below.



Write the Empirical Formula for each compound.  
(Use subscripts to show the number of atoms of each element.)

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_2\text{H}_4\text{O}_2$	
$\text{C}_3\text{H}_6\text{O}_3$	
$\text{C}_2\text{H}_4\text{O}_2$	

Calculate the Molar Mass for the following compounds.

1)  $\text{H}_2\text{O}$

2)  $\text{H}_2\text{O}_2$

3)  $\text{H}_2\text{O}$

4)  $\text{H}_2\text{O}_2$

Calculate the Percent Composition for each element in the following compounds.

1)  $\text{H}_2\text{O}$

2)  $\text{H}_2\text{O}_2$

3)  $\text{H}_2\text{O}$

4)  $\text{H}_2\text{O}_2$

5)  $\text{H}_2\text{O}$

6)  $\text{H}_2\text{O}_2$

7)  $\text{H}_2\text{O}$

8)  $\text{H}_2\text{O}_2$

9)  $\text{H}_2\text{O}$

10)  $\text{H}_2\text{O}_2$

Determine the Empirical and Molecular Formula for the compound given the information below.

1) A compound has the following percent composition: 40.00% Hg, 21.87% Cl, 3.75% N, and 34.38% O. The molar mass for Aqueous is 454.0 g/mol.

a. If you have 454 g of Aqueous, how many grams of Hg, Cl, and O do you have?

b. Determine the Empirical Formula for Aqueous.

c. Determine the Molecular Formula for Aqueous.

# Study Guide The Mole

**Paul B. Kelter**



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