

Name Key  
Date/Period \_\_\_\_\_

## Single Replacement Reaction Worksheet

For each reaction use the activity series to complete the reaction. If no reaction will occur, write "no reaction". Make sure you balance your final answer.

1.  $\text{Ag} + \text{KNO}_3 \rightarrow$  No Reaction
2.  $\text{Zn} + 2\text{AgNO}_3 \rightarrow$   $\text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$
3.  $\text{Cl}_2 + 2\text{KI} \rightarrow$   $2\text{KCl} + \text{I}_2$
4.  $\text{Cu} + \text{FeSO}_4 \rightarrow$  No Reaction
5.  $2\text{Fe} + 3\text{Pb}(\text{NO}_3)_2 \rightarrow$   $2\text{Fe}(\text{NO}_3)_3 + 3\text{Pb}$
6.  $\text{Cu} + \text{Al}_2(\text{SO}_4)_3 \rightarrow$  No Reaction
7.  $2\text{Al} + 3\text{Pb}(\text{NO}_3)_2 \rightarrow$   $2\text{Al}(\text{NO}_3)_3 + 3\text{Pb}$
8.  $\text{Cl}_2 + 2\text{NaI} \rightarrow$   $2\text{NaCl} + \text{I}_2$
9.  $\text{Fe} + 3\text{AgC}_2\text{H}_3\text{O}_2 \rightarrow$   $3\text{Ag} + \text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_3$
10.  $2\text{Al} + 3\text{CuCl}_2 \rightarrow$   $3\text{Cu} + 2\text{AlCl}_3$
11.  $\text{Br}_2 + \text{CaI}_2 \rightarrow$   $\text{CaBr}_2 + \text{I}_2$
12.  $2\text{Fe} + 3\text{CuSO}_4 \rightarrow$   $3\text{Cu} + \text{Fe}_2(\text{SO}_4)_3$
13.  $\text{Cl}_2 + \text{MgI}_2 \rightarrow$   $\text{MgCl}_2 + \text{I}_2$

Activity Series	
Metals	Nonmetals
Li	$\text{F}_2$
Rb	$\text{Cl}_2$
K	$\text{Br}_2$
Cs	$\text{I}_2$
Ba	
Sr	
Ca	
Na	
Mg	
Al	
Ti	
Mn	
Zn	
Cr	
Fe	
Co	
Ni	
Sn	
Pb	
$\text{H}_2$	
Cu	
Ag	
Au	

# Single Replacement Reaction Ws Key

**Alan Sherman**



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