

Concentration Practice Problems

% Solution

1. What is the concentration (% solution) of a solution that contains 0.78 grams of $MgCl_2$ dissolved in 10. grams of water?

$$\% = \frac{\text{mass of solute}}{\text{mass of sol'n}} = \frac{0.78\text{g}}{10.78\text{g}} = 7.24\%$$

2. How many grams of solid are needed to prepare a 10.0% $NaCl$ solution in 200. grams of water?

$$0.10 = \frac{x}{x+200}$$

$$0.10(x+200) = x$$

$$0.10x + 20 = x$$

$$-0.10x \quad -0.10x$$

$$\frac{20}{0.9} = \frac{0.9x}{0.9}$$

$$x = 22.2 \text{ g } NaCl$$

3. What is the % solution of a mixture that contains 0.0500 moles of NaI dissolved in 25.0 g of water?

$$0.0500 \text{ moles } NaI \cdot \frac{150.0\text{g}}{1 \text{ mole}} = 7.5 \text{ g } NaI$$

$$\frac{7.5 \text{ g } NaI}{(25.0\text{g sol'n})} = \frac{29.9\%}{23.1}$$

Molarity

4. What is the molarity of a solution that contains 85.0 g of iodine (I_2) dissolved in 250. mL of ethanol solution?

$$\frac{85.0\text{g}}{1} \cdot \frac{1 \text{ mole}}{254\text{ g}} = 0.335 \text{ mole}$$

$$M = \frac{0.335 \text{ moles}}{0.250 \text{ L}} = 1.34 \text{ M}$$

5. How many grams of $Ca(NO_3)_2$ are need to produce 1500. mL of a 0.75 M solution?

$$M = \frac{\text{mol}}{\text{L}}$$

$$\frac{1.125 \text{ mol}}{1} \cdot \frac{164.1 \text{ g}}{1 \text{ mol}} = 184.6 \text{ g } Ca(NO_3)_2$$

$$0.75 = \frac{x}{1.500\text{L}}$$

$$x = 1.125 \text{ mol}$$

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William F. Hunter

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