

Solution Stoichiometry

1. What volume of a 0.35M AgNO_3 is required to completely react with 25 mL of a 0.24M NaCl solution?



Required: V of AgNO_3

Start with the other reactant

$$\frac{0.24 \text{ mol NaCl}}{1 \text{ L}} \cdot \frac{0.055 \text{ L NaCl}}{1 \text{ mol NaCl}} \cdot \frac{1 \text{ mol AgNO}_3}{1 \text{ mol NaCl}}$$

These cancel

$$\frac{0.35 \text{ mol AgNO}_3}{0.35 \text{ mol AgNO}_3} = 0.037 \text{ L} = 37.7 \text{ mL}$$

2. 37 mL of a 0.24M KI solution is required to completely react with 42 mL of a $\text{Pb}(\text{NO}_3)_2$ solution. What was the original concentration of the lead (II) Nitrate solution?



Given: $\text{V}_1 = 37 \text{ mL KI}$ cm^3
 $\text{V}_2 = 42 \text{ mL Pb}(\text{NO}_3)_2$
 $\text{C}_1 = 0.24 \text{ M}$ KNO_3

Required: original concentration of the lead

$$\frac{0.24 \text{ mol KI}}{1 \text{ L}} \cdot \frac{0.037 \text{ L}}{1} \cdot \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{2 \text{ mol KI}} \cdot \frac{1}{0.037 \text{ L}} = 0.1277$$

$$2(42) \text{ M}_2 = 1(0.24)(37)$$

$$\text{M}_2 = 10.72 / 84$$

$$\text{M}_2 = 0.1277 \text{ M}$$

3. What mass of solid Magnesium Hydroxide can be produced if 15 mL of a 0.63M $\text{Mg}(\text{NO}_3)_2$ solution reacts completely with excess NaOH ?



Required: mass of solid Magnesium Hydroxide if 15 mL of a 0.63M $\text{Mg}(\text{NO}_3)_2$

Given: 15 mL $\text{Mg}(\text{NO}_3)_2$, 0.63M $\text{Mg}(\text{NO}_3)_2$

1 mol AgNO_3 solution:

$$\frac{0.63 \text{ mol Mg}(\text{NO}_3)_2}{1 \text{ L}} \cdot \frac{0.015 \text{ L}}{1} \cdot \frac{1 \text{ mol Mg(OH)}_2}{1 \text{ mol Mg}(\text{NO}_3)_2}$$

$$\cdot \frac{58.32 \text{ g}}{1 \text{ mol Mg(OH)}_2} = 1.65 \text{ g Mg(OH)}_2$$

4. What mass of solid Calcium Phosphate can be produced if 78mL of a 0.17M Sodium Phosphate solution is mixed with 38mL of a 0.21M Calcium Chloride solution?



Given:

$$\frac{78 \text{ mL Na}_3\text{PO}_4}{0.17 \text{ M Na}_3\text{PO}_4} \cdot \frac{38 \text{ mL CaCl}_2}{0.21 \text{ M CaCl}_2}$$

molar mass:

$$\text{Ca} = 2 \cdot 40.07 \quad \text{P} = 2 \cdot 30.97 \quad \text{O} = 2 \cdot 16 \quad \text{Na} = 23$$

Required:
 $\text{mass(g)} \text{ Ca}_3(\text{PO}_4)_2$

From: Na_3PO_4
 $\frac{0.17 \text{ mol Na}_3\text{PO}_4}{1 \text{ L}} \cdot \frac{0.078 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol Na}_3\text{PO}_4}$

$$\cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 1.925 \text{ g Ca}_3(\text{PO}_4)_2$$

From CaCl_2

$$\frac{0.21 \text{ mol CaCl}_2}{1 \text{ L}} \cdot \frac{0.038 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol CaCl}_2} \cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 1.4371 \text{ g Ca}_3(\text{PO}_4)_2$$

Solution Stoichiometry Practice Problems

Valentine Korah

Solution Stoichiometry Practice Problems:

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all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking *The Software Encyclopedia 2000* Bowker Editorial Staff,2000-05 **Introduction to Chemistry, Laboratory Manual** T. R. Dickson,1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students

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