

Balancing Single Replacement Reactions

In a **single replacement** reaction, the cation (usually a metal) in a compound is replaced by another, more reactive metal. They always follow this format:



As an example, look at this chemical equation for single replacement reaction between solid lithium and a solution of iron (II) nitrate.



Lithium replaced iron (II) in the compound. Iron was left alone as a solid product. This reaction, however, is not balanced. Count the atoms. You might find it helpful to treat the nitrate polyatomic ion as a single element in balancing.

Reactants

Li = _____

Fe = _____

NO₃ = _____

Products

Li = _____

Fe = _____

NO₃ = _____

Add coefficients to balance the equation and satisfy the Law of Conservation of Matter.



Not all combinations of elements and compounds in single replacement reactions will actually react! You can use the **activity series** of metals to predict this. Follow these two rules:

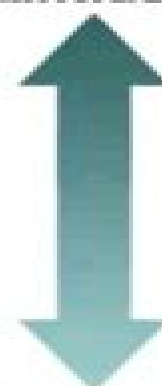
1. *Highly reactive metals prefer to be part of a compound.*
2. *Highly stable metals prefer to be alone.*

Look at the following combinations. If they will react, then predict what the products will be. If they will not react, then write NR for "no reaction".



lithium (Li)
potassium (K)
barium (Ba)
calcium (Ca)
sodium (Na)
magnesium (Mg)
aluminum (Al)
zinc (Zn)
cobalt (Co)
nickel (Ni)
lead (Pb)
copper (Cu)
silver (Ag)
gold (Au)

most reactive



least reactive

Single Replacement Reactions Study Guide

Leo J. Malone, Theodore O. Dolter



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