

Balancing Equations Worksheet

- 1) $\underline{\quad 1 \quad} \text{Na}_3\text{PO}_4 + \underline{\quad 3 \quad} \text{KOH} \rightarrow \underline{\quad 3 \quad} \text{NaOH} + \underline{\quad 1 \quad} \text{K}_3\text{PO}_4$
- 2) $\underline{\quad 1 \quad} \text{MgF}_2 + \underline{\quad 1 \quad} \text{Li}_2\text{CO}_3 \rightarrow \underline{\quad 1 \quad} \text{MgCO}_3 + \underline{\quad 2 \quad} \text{LiF}$
- 3) $\underline{\quad 1 \quad} \text{P}_4 + \underline{\quad 3 \quad} \text{O}_2 \rightarrow \underline{\quad 1 \quad} \text{P}_2\text{O}_3$
- 4) $\underline{\quad 2 \quad} \text{RbNO}_3 + \underline{\quad 1 \quad} \text{BeF}_2 \rightarrow \underline{\quad 1 \quad} \text{Be}(\text{NO}_3)_2 + \underline{\quad 2 \quad} \text{RbF}$
- 5) $\underline{\quad 2 \quad} \text{AgNO}_3 + \underline{\quad 1 \quad} \text{Cu} \rightarrow \underline{\quad 1 \quad} \text{Cu}(\text{NO}_3)_2 + \underline{\quad 2 \quad} \text{Ag}$
- 6) $\underline{\quad 1 \quad} \text{CF}_4 + \underline{\quad 2 \quad} \text{Br}_2 \rightarrow \underline{\quad 1 \quad} \text{CBr}_4 + \underline{\quad 2 \quad} \text{F}_2$
- 7) $\underline{\quad 2 \quad} \text{HCN} + \underline{\quad 1 \quad} \text{CuSO}_4 \rightarrow \underline{\quad 1 \quad} \text{H}_2\text{SO}_4 + \underline{\quad 1 \quad} \text{Cu}(\text{CN})_2$
- 8) $\underline{\quad 1 \quad} \text{GaF}_3 + \underline{\quad 3 \quad} \text{Cs} \rightarrow \underline{\quad 3 \quad} \text{CsF} + \underline{\quad 1 \quad} \text{Ga}$
- 9) $\underline{\quad 1 \quad} \text{BaS} + \underline{\quad 1 \quad} \text{PtF}_2 \rightarrow \underline{\quad 1 \quad} \text{BaF}_2 + \underline{\quad 1 \quad} \text{PtS}$
- 10) $\underline{\quad 1 \quad} \text{N}_2 + \underline{\quad 3 \quad} \text{H}_2 \rightarrow \underline{\quad 2 \quad} \text{NH}_3$
- 11) $\underline{\quad 2 \quad} \text{NaF} + \underline{\quad 1 \quad} \text{Br}_2 \rightarrow \underline{\quad 2 \quad} \text{NaBr} + \underline{\quad 1 \quad} \text{F}_2$
- 12) $\underline{\quad 1 \quad} \text{Pb}(\text{OH})_2 + \underline{\quad 2 \quad} \text{HCl} \rightarrow \underline{\quad 2 \quad} \text{H}_2\text{O} + \underline{\quad 1 \quad} \text{PbCl}_2$
- 13) $\underline{\quad 2 \quad} \text{AlBr}_3 + \underline{\quad 3 \quad} \text{K}_2\text{SO}_4 \rightarrow \underline{\quad 6 \quad} \text{KBr} + \underline{\quad 1 \quad} \text{Al}_2(\text{SO}_4)_3$
- 14) $\underline{\quad 1 \quad} \text{CH}_4 + \underline{\quad 2 \quad} \text{O}_2 \rightarrow \underline{\quad 1 \quad} \text{CO}_2 + \underline{\quad 2 \quad} \text{H}_2\text{O}$
- 15) $\underline{\quad 2 \quad} \text{Na}_3\text{PO}_4 + \underline{\quad 3 \quad} \text{CaCl}_2 \rightarrow \underline{\quad 6 \quad} \text{NaCl} + \underline{\quad 1 \quad} \text{Ca}_3(\text{PO}_4)_2$
- 16) $\underline{\quad 2 \quad} \text{K} + \underline{\quad 1 \quad} \text{Cl}_2 \rightarrow \underline{\quad 2 \quad} \text{KCl}$
- 17) $\underline{\quad 2 \quad} \text{Al} + \underline{\quad 6 \quad} \text{HCl} \rightarrow \underline{\quad 3 \quad} \text{H}_2 + \underline{\quad 2 \quad} \text{AlCl}_3$
- 18) $\underline{\quad 1 \quad} \text{N}_2 + \underline{\quad 3 \quad} \text{F}_2 \rightarrow \underline{\quad 2 \quad} \text{NF}_3$
- 19) $\underline{\quad 1 \quad} \text{SO}_2 + \underline{\quad 2 \quad} \text{Li}_2\text{Se} \rightarrow \underline{\quad 1 \quad} \text{SSe}_2 + \underline{\quad 2 \quad} \text{Li}_2\text{O}$
- 20) $\underline{\quad 2 \quad} \text{NH}_3 + \underline{\quad 1 \quad} \text{H}_2\text{SO}_4 \rightarrow \underline{\quad 1 \quad} (\text{NH}_4)_2\text{SO}_4$

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Disha Experts

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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1. Understanding the eBook Skill Practice 51 Answers Chemistry
 - The Rise of Digital Reading Skill Practice 51 Answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 51 Answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 51 Answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 51 Answers Chemistry
 - Personalized Recommendations
 - Skill Practice 51 Answers Chemistry User Reviews and Ratings
 - Skill Practice 51 Answers Chemistry and Bestseller Lists
5. Accessing Skill Practice 51 Answers Chemistry Free and Paid eBooks
 - Skill Practice 51 Answers Chemistry Public Domain eBooks
 - Skill Practice 51 Answers Chemistry eBook Subscription Services
 - Skill Practice 51 Answers Chemistry Budget-Friendly Options
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- ePub, PDF, MOBI, and More
- Skill Practice 51 Answers Chemistry Compatibility with Devices
- Skill Practice 51 Answers Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 51 Answers Chemistry
 - Highlighting and Note-Taking Skill Practice 51 Answers Chemistry
 - Interactive Elements Skill Practice 51 Answers Chemistry
- 8. Staying Engaged with Skill Practice 51 Answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 51 Answers Chemistry
- 9. Balancing eBooks and Physical Books Skill Practice 51 Answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 51 Answers Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 51 Answers Chemistry
 - Setting Reading Goals Skill Practice 51 Answers Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 51 Answers Chemistry
 - Fact-Checking eBook Content of Skill Practice 51 Answers Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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