

NAME: \_\_\_\_\_

# Molarity Practice

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1. Calculate the molarity of the following solutions:

a) 45.0 g of  $\text{NaCl}$  in 2.00 L of solution.b) 10.0 g of  $\text{KNO}_3$  in 500 mL of solution.c) 175.0 g of  $\text{CuSO}_4$  in 2.50 L of solution.

2. What is the molarity of a solution that contains 2.0 g of each of the following?

a)  $\text{NaCl}$  b)  $\text{KNO}_3$ c)  $\text{H}_2\text{SO}_4$  d)  $\text{NaOH}$ 

3. What is the molarity of a solution that contains 10.0 g of each of the following in 2.00 L of solution?

4. Which of the following contains the highest concentration of ions per liter of solution?

a) 10.0 g of  $\text{NaCl}$  in 1.00 L of solutionb) 20.0 g of  $\text{NaCl}$  in 1.00 L of solution

5. Make a more concentrated solution of sodium chloride (NaCl) and, at the same time, make a solution that has a concentration of 1.0 M. What are some ways to make a more concentrated solution? (You may use the molarity and the mass, or find the molar mass and molarity.)

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