

# Concentration of Solutions

## Teacher Notes and Answers

### Exercise 1: Concentration of Solutions

- Concentration is a measure of the amount of substance (solute) in a solution, and is expressed in a number of different ways. The most common is molarity, which is the amount of solute in a given volume of solution.
- The molarity of a solution is calculated by dividing the amount of solute (in moles) by the volume of the solution (in litres).
- To calculate the concentration of a solution, you need to know the amount of solute and the volume of the solution. The concentration is then calculated by dividing the amount of solute by the volume of the solution.

### Exercise 2

- 0.1 mol/L
- 0.2 mol/L
- 0.3 mol/L
- 0.4 mol/L
- 0.5 mol/L
- 0.6 mol/L
- 0.7 mol/L
- 0.8 mol/L
- 0.9 mol/L
- 1.0 mol/L

### Exercise 3

- 0.1 mol/L
- 0.2 mol/L
- 0.3 mol/L
- 0.4 mol/L
- 0.5 mol/L
- 0.6 mol/L
- 0.7 mol/L
- 0.8 mol/L
- 0.9 mol/L
- 1.0 mol/L

# Study Guide Solution Concentration Answers

**Michael Brown**



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