

The Mole Study Guide

Mass Composition: What's your element's %? given after the decimal? Molar mass to obtain your value!

10. There are 3.75 moles of CO_2 .

11. There are _____ moles of _____.

12. There are _____ moles of _____.

13. There are _____ moles of _____.

14. There are 2.5 moles of _____.

15. There are _____ moles of _____.

16. There are _____ moles of _____.

17. There are _____ moles of _____.

18. There are 4.5 moles of _____.

19. There are _____ moles of _____.

20. There are _____ moles of _____.

21. There are 3.5 moles of _____.

22. There are _____ moles of _____.

23. There are _____ moles of _____.

24. There are 2.5 moles of _____.

25. There are _____ moles of _____.

26. There are _____ moles of _____.

Name _____ Period _____ Date _____

The Mole Study Guide

Fill in the Blank boxes for the image below.



Write the Empirical Formula for each compound.
(Use subscripts to show the relative ratio of atoms.)

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_2\text{H}_4\text{O}_2$	
$\text{C}_3\text{H}_6\text{O}_3$	
$\text{C}_2\text{H}_4\text{O}_2$	

Calculate the Molar Mass for the following compounds.

10. H_2O

11. NaCl

12. H_2SO_4

13. NaCl

Calculate the Percent Composition for each element in the following compound.

10. H_2O

11. NaCl

12. H_2O

13. NaCl

14. H_2O

15. NaCl

16. H_2O

17. NaCl

18. H_2O

19. NaCl

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 40.00% Hg, 21.87% Cl, 3.12% N, and 34.97% O. The molar mass for Astatine is 210.0 g/mol.

a. If you have 100 g of Astatine, how many grams of Hg, Cl, and O do you have?

b. Determine the Empirical Formula for Astatine.

c. Determine the Molecular Formula for Astatine.

Study Guide The Mole

**Katherine J. Denniston, Denniston
Katherine**



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